

Synthesis, structural characterization and catalytic application of citrate-stabilized monometallic and bimetallic palladium@copper nanoparticles in microbial anti-activities

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Abstract: In this research work, copper (Cu), palladium (Pd) and their bimetallic palladium@copper (Pd@Cu) nanoparticles were synthesized using trisodium citrate as a stabilizing agent using the known chemical reduction method. The synthesized Cu, Pd and Pd@Cu nanoparticles were characterized by the ultraviolet-visible spectroscopy, scanning electron microscopy and X-ray diffraction spectroscopy, respectively. The different volumes of trisodium citrate were used for the stability of synthesized monometallic Cu, Pd and bimetallic Pd@Cu nanoparticles. The synthesized Cu, Pd and their bimetallic Pd@Cu nanoparticles were used as catalysts for the reduction of 4-nitrophenol in the presence of NaBH₄. The bimetallic Pd@Cu nanoparticles had efficient catalytic activities with a high rate constant (1.812 min⁻¹) as compared to monometallic Cu (0.3322 min⁻¹) and Pd (0.2689 min⁻¹) nanoparticles, respectively. The correlation coefficient (*R*²) was found to be 0.99 for these three nanoparticles. Meanwhile, the effect of Cu, Pd and bimetallic Pd@Cu nanoparticles was checked on the physiology of specific different micro-organism strains. The bimetallic Pd@Cu nanoparticles reported the maximum resistance at maximum level the growth of bacterial strain and had observed a smooth antibacterial graph than the monometallic analogs.

Keywords: stabilizing agent, bimetallic nanoparticles, catalytic activities, antimicrobial effects

Introduction

Nanotechnology deals with small chemical materials ranging from 1 to 100 nm, having specific physicochemical properties that are totally different from their bulk materials.¹⁻³ In the last few decades, nanotechnology has been gaining a tremendous reputation in the field of biotechnology, electronics, drug delivery, catalysis, and so on.^{4,5} Nanomaterials have size- and shape-dependent properties, which make them totally different from their bulk counterparts, while the bulk materials have size- and shape-independent properties.⁶ Due to the surface plasmon resonance, metal nanoparticles show different colors by changing shape and particle size.⁷ The nanoparticles with smaller size than the wavelength of electromagnetic radiations and having high penetrating power are efficiently used as conducting materials.⁸ Similarly, the melting point of this material is automatically decreased with decrease in particle size, probably up to <10 nm.⁶ Metal nanoparticles are very attractive because of its high mechanical strength, thermal conductivity, electrical conductivity, chemical stability and large surface area.⁹ Current advancements in nanomaterials have guaranteed

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that the metal nanoparticles can play an important role in the different concerned technologies.^{10–12}

Nowadays, metal-fabricated nanomaterials have been gaining extensive interest because of their enriched optical, catalytic and antibacterial properties.^{13,14} Hence, it is made easier to recover and reuse the catalysts by incorporation of magnetic nanoparticles in catalysts or absorbents simply by applying the external magnetic field.¹⁵ That is why, different metal nanoparticles, particularly the gold, cobalt, nickel and silver, have been synthesized.^{16,17} Among these metals, gold and silver nanoparticles are much used in different fields because of their conductivity, inertness, surface plasmon resonance and bright color.¹⁸ However, the application of gold at industrial scale is restricted because of its high cost. Nevertheless, copper (Cu) nanoparticles are extensively used because of their low cost in good electrical, thermal, catalytic and biomedical applications.¹⁹ Moreover, Cu nanoparticles can act as antibiotics and antimicrobial and antifungal agents particularly in the plastic, coating and textile industries.^{20,21} Similarly, palladium (Pd) nanoparticles have been widely studied in a wide range of catalytic applications including hydrogenations, oxidations, carbon–carbon bond formation, and electrochemical reactions in the fuel cells.^{22,23} However, it should be noted that the application of Pd goes beyond catalysis. For example, the property of Pd to adsorb hydrogen has also led to use of Pd-based nanoparticles in hydrogen storage and sensing applications. There are many types of environmental pollutants, Congo red, methylene blue, Rhodamine B, Rhodamine G, methyl orange, eosin Y and nitro compounds.²⁴ The compounds having nitro group are selected on the basis of their utmost importance in environment and in various manufacturing industries in general, particularly in the pharmacology products, aromatic products, dyes and pigments. Mostly, in pharmaceutical sector, the antipyretic and analgesic type of medicines, that is, paracetamol and acetanilide etcare, are synthesized from the 4-aminophenol, which are produced by the catalytic activities of 4-nitrophenol (4-NP). Similarly, the aforementioned compounds were prepared by the reduction of 2 nitrophenol (2-NP) and 4-nitroaminophenol (4-NA), respectively. These compounds are mostly used as a source in industry for the production of different pigments, dyes, and so on. Although, the nitro-based compounds have categorized in pollutants and toxic materials which are discharge by synthetic industries after process completion, in terms of waste materials. That is why, its reduction and conversion in terms of sustainability to useful and worthy compounds are very important and the utmost needs of recent world.²⁵

Since last few decades, metallic nanocomposites are gaining much more attention because of their efficient catalytic,

opto-electronic and antibacterial properties.^{26,27} Although, their simple synthesis via chemical route, that is, concurrent and successive reduction of the precursor ions, gives various morphological composite nanoparticles. However, the nanoparticles have a wide surface for energy and the presence of van der Waal forces and aggregate it readily, which leads to decrease in the specific surface area that comprehensively decreases the catalytic activity. Stabilizers in the form of polymers, surfactants or any other legends are added during reduction reaction in order to protect nanosize material from aggregation.²⁸ Among the following types of bimetallic nanoparticles, Pd-containing bimetallic nanoparticles like Pd@Ni, Pd@Au, Pd@Pt and Pd@Cu are considered as promising candidates for various organic transformation reactions.²⁹ Moreover, it has been experimentally investigated that the catalytic properties of bimetallic nanoparticles are often affected by a range of factors, like nanostructure, morphological composition of particle, size and shape. The important features that distinguish the bimetallic catalyst from monometallic ones are the tunable nanostructures and surface compositions. The bimetallic nanostructure may eventually be alloy, core-shell structure or mixed monometallic nanoparticles. These different nanostructures establish different catalytic performances.³⁰

In the present research work, Cu, Pd and the bimetallic Pd@Cu nanoparticles were synthesized through a known chemical reduction method using sodium borohydride (NaBH_4) as a reducing agent and trisodium citrate as a stabilizer. The synthesized monometallic and bimetallic nanoparticles were checked for its catalytic and antimicrobial activities. Although many researchers have reported the synthesis and characterization of Cu and Pd nanoparticles, our literature survey shows deficiencies with respect to their antimicrobial activities and catalytic properties as well as in terms of toxic environmental pollutants degradation.

Experimental section

Materials

Copper nitrate trihydrates ($\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$, 99%), palladium chloride (PdCl_2) and trisodium citrate dihydrate ($\text{C}_6\text{H}_5\text{O}_7\text{Na}_3 \cdot 2\text{H}_2\text{O}$, 99%) were purchased from Sigma-Aldrich company and used as received without any amendment and purification. Sodium borohydride (NaBH_4 , 98%) was used as a reducing agent for the synthesis of nanoparticles and reduction of 4-NP (99%). The NaOH (99%), L-ascorbic acid, liquid Mueller Hinton, Luria broth (LB) and ofloxacin were purchased from Merck and Aldrich and used as received. The ultrapurified water was obtained from Milli-Q ultrapure (18.2 MN cm) system and was used in all

the experimentation setup. The different bacterial species strains used were *Proteus mirabilis*, *Bacillus thuringiensis*, *Shigella flexneri*, *Staphylococcus aureus*, *Klebsiella pneumoniae*, *Escherichia coli*, *Pseudomonas aeruginosa* and *Salmonella typhimurium*.

Synthesis of Cu, Pd and Pd@Cu bimetallic nanoparticles

Copper nanoparticles were synthesized by the following procedure as reported by Fang et al with slight modification.³¹ 0.5 mL aqueous solution of trisodium citrate (0.4 M) and 5.0 mL aqueous solution of $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ (0.5 M) were added in a flask containing 30 mL deionized (DI) water and stirred constantly. After 30 min, 0.5 mL freshly prepared aqueous solution of NaBH_4 (0.1 M) was added to the mixture. The reaction was continued till their color turned deep red, indicating the synthesis of Cu nanoparticles. The same procedure was followed up for the synthesis of Pd nanoparticles by adding 0.5 mL aqueous solution of trisodium citrate (0.4 M) and 5 mL aqueous solution of PdCl_2 (0.10 M) in a flask containing 30 mL DI water, and stirred continuously. After 30 min, 0.5 mL of the already prepared aqueous solution of NaBH_4 (0.1 M) was added into the flask. The reaction was continued under constant stirring, until the color turned to dark brown, which indicated the successful synthesis of Pd nanoparticles.

Moreover, Pd nanoparticles were deposited over Cu nanoparticles to synthesize the bimetallic Pd@Cu nanoparticles. 0.5 mL L-ascorbic acid (0.2 M) was added to 15 mL prepared Pd nanoparticles (0.5 M) in a given flask with a constant stirring rate. After 30 min, 15 mL of $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ solution (0.05 M) was added to the reaction in the presence of nitrogen gas. After 30 min, 2.5 mL trisodium citrate (0.5%) was added dropwise to the reaction mixture. In the end, 2 mL NaOH (0.6 M) was added to the mixture dropwise to adjust the pH at 8.5 for 2 h at room temperature, to produce the bimetallic Pd@Cu nanoparticles. The resulting monometallic and bimetallic nanoparticles were centrifuged at 1,800 rpm for 20 min. The proposed compound was gently washed and redispersed in 10 mL DI water.

Structural characterization of Cu, Pd and Pd@Cu bimetallic nanoparticles

Ultraviolet–visible (UV-Vis) spectroscopy was performed using Shimadzu UV-2450 spectrometer equipped with temperature-controlled program (TCC-240A) to study the absorbance maxima and catalytic activities of synthesized monometallic and bimetallic nanoparticles. The phase purity of bimetallic Pd@Cu nanoparticles was analyzed

by X-ray diffraction (XRD) using Bruker D2 PHASER X-ray diffractometer with graphite monochromator using Cu $\text{K}\alpha$ radiation ($\lambda=1.54056 \text{ \AA}$), operating at 30 kV and 15 mA. Similarly, Pd@Cu nanoparticles were characterized by energy-dispersive X-ray spectroscopy (EDS). The morphology of synthesized monometallic and bimetallic nanoparticles was studied by scanning electron microscopy (JSM 6700F). A sample was dropped on a silicon single crystal sheet, dried by infrared light and carefully placed on conducting glue. The sample was coated with gold vapor in order to make conduction for analysis by using JSM 6700F scanning electron microscope.

Catalytic activities of Cu, Pd and Pd@Cu bimetallic nanoparticles

Checking out the catalytic activity of synthesized Cu, Pd and Cu@Pd bimetallic catalysts during the reduction of 4-NP to 4-aminophenol by NaBH_4 was investigated as a model reaction at room temperature. During the reductive degradation of 4-NP in the presence of NaBH_4 , the following typical procedure was followed. Aqueous solution of each catalyst (0.1 mg mL^{-1}) and NaBH_4 (0.1 M) was prepared. Forty microliters of aqueous solution of 4-NP, 150 μL aqueous solution of NaBH_4 and 3 mL DI water were added in UV quartz cell followed by adding 15 μL of synthesized Cu, Pd and Pd@Cu bimetallic catalysts. The progressive sequences of catalysts were investigated by changing observation in the mixture reaction at specific interval within the range of 250–500 nm under UV-Vis spectrophotometer.

During the recyclization experiment, the catalyst was isolated via centrifugation. After centrifugation, the following procedure was used to recover and reuse the catalyst. The catalyst was washed 3 times with DI water and dispersed in 10 μL and reused in the subsequent reaction. The nitro compounds as 4-NP is a typical example of toxic and persistent pollutant present in industrial waste water, soil and effluents that have an adverse effect directly on the aquatic life and human health, in general. The common substance that is regulated by USA Environmental Protection Agency is 4-NP. It is considered as a model for catalytic degradation studies. The catalytic properties of citrate-stabilized Cu and Pd nanoparticles were studied for the reduction of 4-NP in the presence of excess amount of NaBH_4 in a quartz cell using UV-Vis spectroscopy.

During the process, 1 mL aqueous solution of 4-NP (0.1 mM) was mixed with 2 mL aqueous solution of NaBH_4 (20 mM). The mixture was adjusted to the desired temperature at 30°C and poured into quartz cell. The catalyst temperature (Cu and Pd nanoparticles) was adjusted to the

same temperature. After adjusting the proposed temperature, 0.5 mL of (0.1 mg mL⁻¹) each catalyst was separately added into 4-NP solution in the quartz cell to record the decrease in absorbance maxima due to reduction of 4-NP into 4-aminophenol. The reduction rate of 4-NP was determined by calculating the decrease in absorbance maxima at 400 nm, accordingly. In case of bimetallic Pd@Cu nanoparticles, the catalytic properties were studied by using the same conditions as adjusted for monometallic Pd and Cu, respectively. The temperature of mixture and bimetallic nanoparticle catalyst was separately adjusted to 30°C by using water bath with controlled temperature. The reduction rate of 4-NP was determined from the decrease in absorption maxima at 400 nm, within the interval time.

Antibacterial liability assessment of Cu, Pd and Pd@Cu bimetallic nanoparticles

The antibacterial properties of prepared Cu, Pd and Pd@Cu bimetallic nanoparticles were investigated against the Gram⁺ and Gram⁻ bacteria using common quantitative agar disk method.³² The Gram⁺ and Gram⁻ bacteria culture was cultivated on LB agar media (20 mL) with pH 7.3 at 25°C. A microbial culture (0.1 mL) suspended in broth containing probably 10⁻⁶ colony-forming units per milliliter was spread uniformly over the surface of agar plate. The prepared Cu, Pd and their bimetallic Pd@Cu nanoparticles (15 µg)/mL of each catalyst were separately suspended in sterilized distilled water and a blank sterilized Whitman filter paper (No 1, diameter 4 mm) was soaked in the nanoparticle suspensions.

The nanoparticle-loaded disk was put over the agar plate by using sterilized forceps.

The above experiment was repeated three times consecutively and the diameters of inhibition zones were calculated after 24 h at 37°C. The antibacterial drug used as a control during the experiments was ofloxacin and the diameters of inhibition zones were measured in millimeters.

Results and discussion

Morphology and size of Cu, Pd and their bimetallic Pd@Cu nanoparticles

The morphology of citrate-stabilized monometallic Cu, Pd and their bimetallic Pd@Cu nanoparticles was investigated using scanning electron micrographs (SEMs) technique. The SEM images of Cu and Pd nanoparticles showed monodispersity and spherical shape as shown in Figure 1A and B corresponding to their size distribution plots. The average size of Cu and Pd nanoparticles determined by nanomeasure software were adjusted as 40 and 50 nm, respectively. The particle sizes were determined by nanomeasure software.³³

Similarly, the morphology and size distribution of bimetallic Pd@Cu nanoparticles were studied using SEM as shown in Figure 1C with the corresponding size distribution plots. It was observed that the average size of bimetallic Pd@Cu nanoparticles measured by nanomeasure software was 380 nm (0.38 µm). Meanwhile, XRD analysis of the bimetallic Pd@Cu nanoparticles was carried out as shown in Figure 2. The composition and structural verification of the prepared Pd@Cu bimetallic nanoparticles by using the XRD

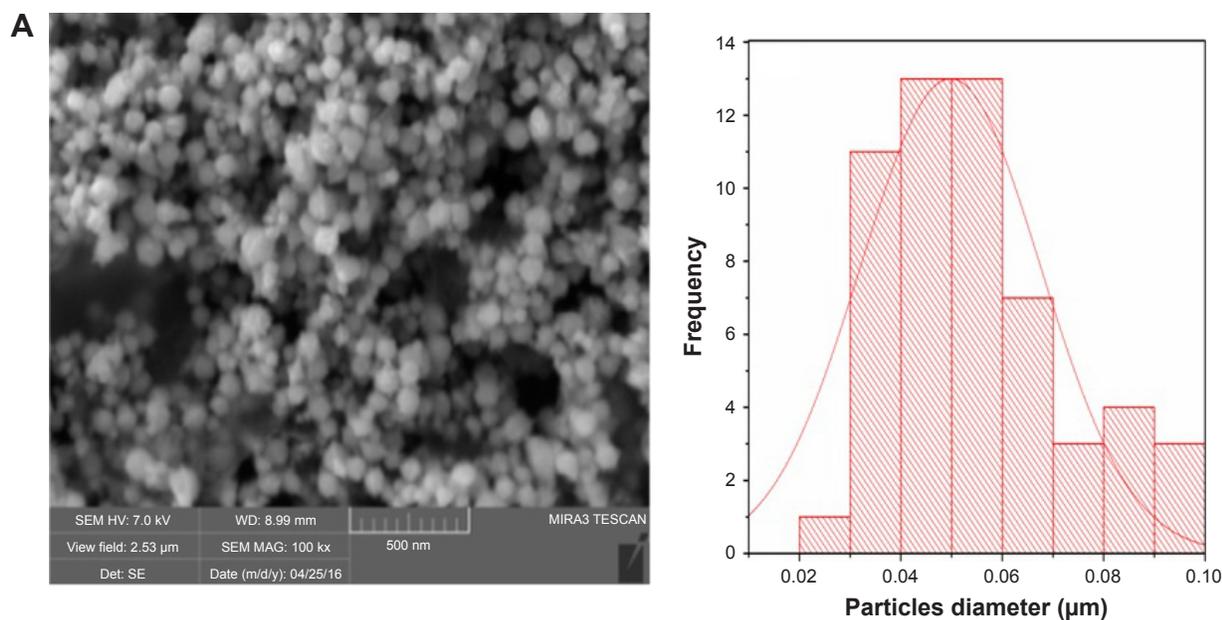


Figure 1 (Continued)

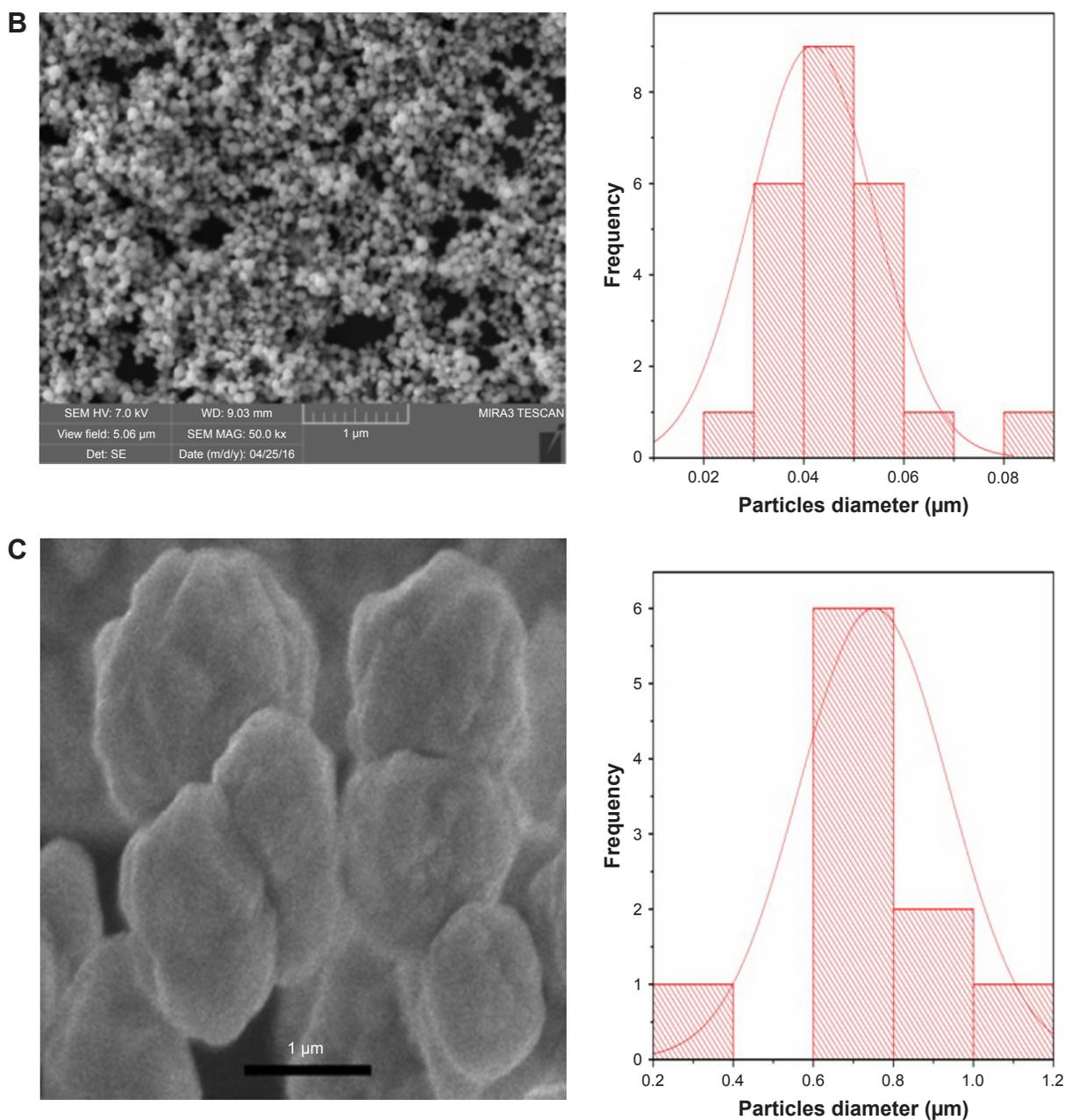


Figure 1 Scanning electron micrographs and size distribution plots of (A) copper, (B) palladium and (C) palladium@copper bimetallic nanoparticles.

and energy-dispersive X-ray are shown in Figure 2A and B, respectively. The XRD peaks caused by Cu-Pd bimetallic nanoparticles indexed as a face-centered cubic structure among 2/3 lie on the pure face-centered cubic Pd (JCPDS no 46–1043).

Meanwhile, lattice parameters of Cu@Pd bimetallic nanoparticle constituents suggest the possibility of alloy structure. A thin film of bimetallic Pd@Cu nanoparticle sample exhibited diffraction peaks at 41, 48, 70, 85 and 90° corresponding to the crystal surfaces 111, 200, 220, 311 and 222, respectively, compared to the standard diffraction

pattern.³⁴ The confirmation of each nanoparticle, that is, Pd, Cu and alloy, formation was proved by energy-dispersive X-ray elemental mapping analysis as shown in Figure 2 (B1–3). The Pd@Cu bimetallic nanoparticle composition was proved by using the EDS and the Pd and Cu detected the peaks of Mo- and carbon-coated SEM grid as shown in Figure S1.

The three types of metal nanoparticles, that is, Cu, Pd and their bimetallic nanoparticles, were prepared by using different volumes (0.2–0.6 mL) of trisodium citrate. The effect of change in amount of trisodium citrate was studied by using

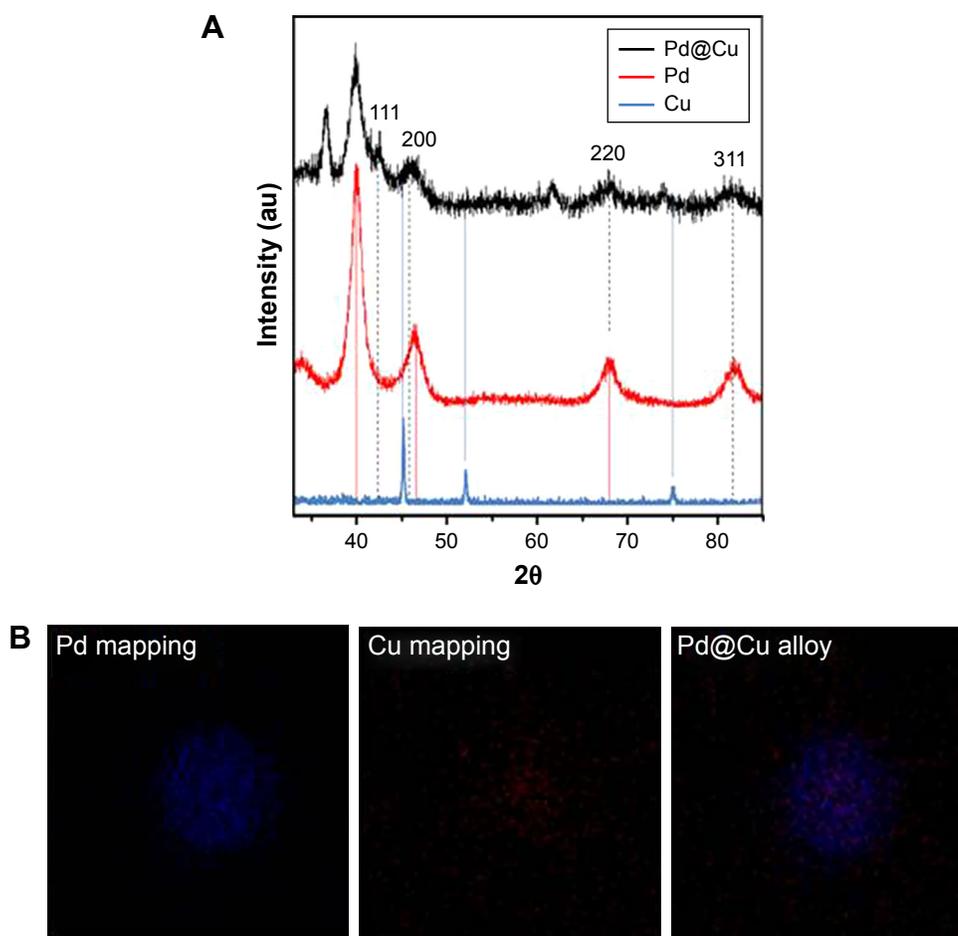


Figure 2 (A) X-ray diffraction pattern of palladium@copper bimetallic nanoparticles. **(B)** Energy-dispersive X-ray elemental mapping images of Pd, Cu, and Pd@Cu bimetallic nanoparticles, respectively.

UV-Vis spectroscopy and the relative change in absorbance maxima for Cu, Pd and Pd@Cu bimetallic nanoparticles as shown in Figure 3A–C. The absorbance maxima appeared at 600 and 400 nm for Cu and Pd nanoparticles for the different volumes of trisodium citrate. Similar results were obtained for bimetallic Pd@Cu nanoparticles and their absorbance maxima appeared at 415 nm for different volumes of trisodium citrate. The appearance of absorbance maxima for Pd@Cu nanoparticles at 415 nm showed good deposition of Pd nanoparticles over Cu nanoparticles, and the absorbance maxima of bimetallic nanoparticles were very close to the absorbance maxima of Pd nanoparticles. Moreover, the stability of absorbance maxima for the volume of trisodium citrate indicated the synthesis of same-size Cu, Pd and bimetallic Pd@Cu nanoparticles,⁷ whereas the increase in absorption peak intensity showed increase in the synthesis of nanoparticles per unit area in medium with increase in volume of trisodium citrate.

Catalytic activities of Cu, Pd and their bimetallic Pd@Cu nanoparticles

The catalytic properties of prepared monometallic Cu, Pd and their bimetallic Pd@Cu nanoparticles were investigated by reducing 4-NP to 4-aminophenol in the presence of excess amount of NaBH₄ as a reducing agent. The reduction process was monitored by UV-Vis spectroscopy at regular intervals. After the addition of NaBH₄ to aqueous solution of 4-NP, the color of the solution changed from light yellow to deep yellow because of conversion of 4-NP to 4-nitrophenolate ion as shown in Figure 4A. Similarly, the absorption maxima of 4-NP changed from 318 to 400 nm because of conversion of 4-NP into 4-nitrophenolate ion in the presence of NaBH₄, and is presented in Figure 4B.³⁵ It was also investigated that neither metal nanoparticles nor NaBH₄ alone can degrade the 4-NP⁵. However, by the addition of metal nanoparticles, that is, Cu, Pd and Pd@Cu bimetallic nanoparticles, into an aqueous solution of 4-NP in the presence of NaBH₄,

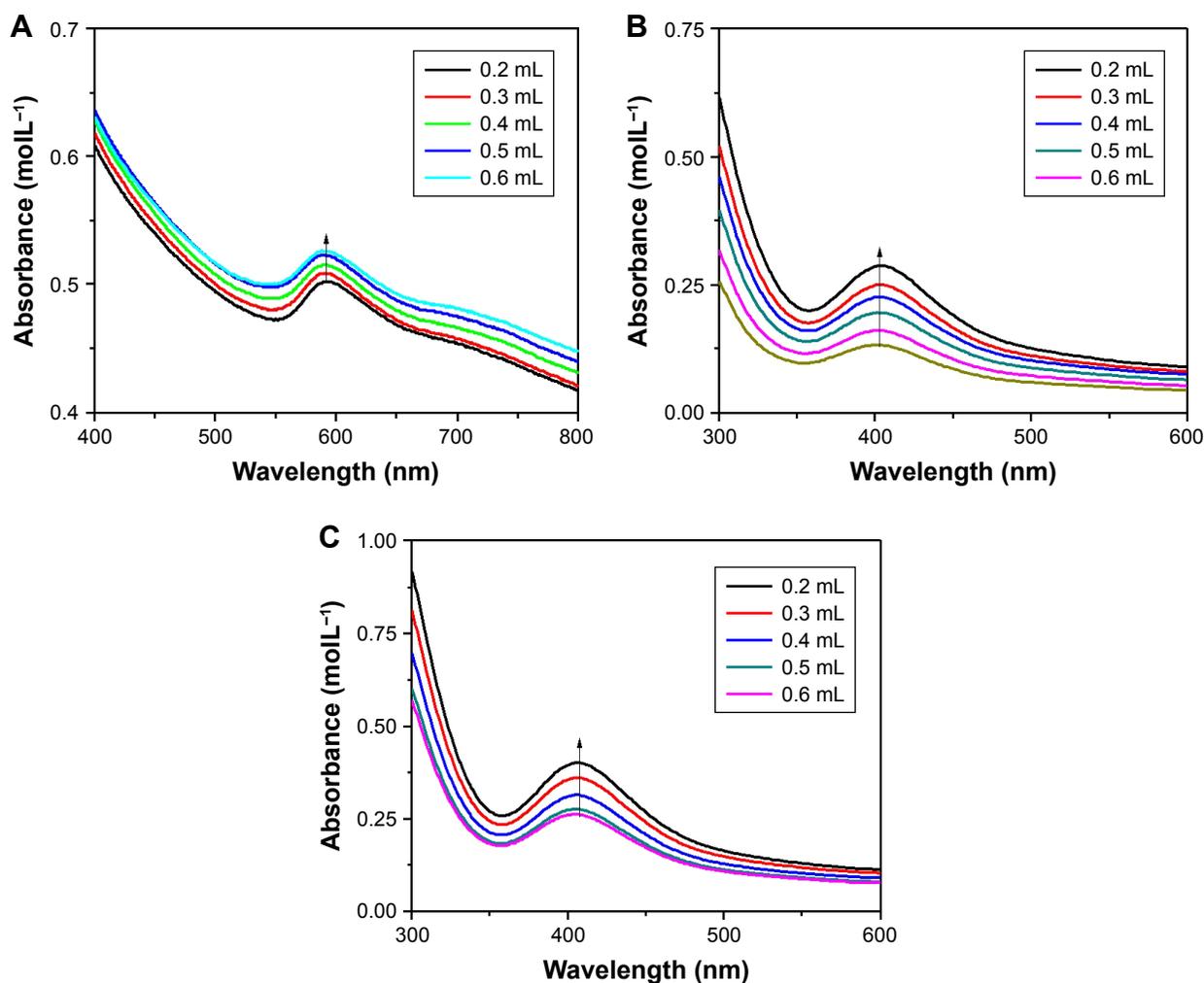


Figure 3 Absorbance spectra of (A) copper, (B) palladium and (C) palladium@copper bimetallic nanoparticles using different volumes of trisodium citrate (5 mM). The arrows show the maximum absorbance.

a sharp decrease in absorption maxima was observed for 4-nitrophenolate ions as shown in Figure 5A–C. The peak that appeared at 300 nm is assigned to the formation of 4-aminophenols.³⁶ The apparent rate (k_{app}) constant and correlation constants (R^2) for Cu, Pd and Pd@Cu nanoparticles were calculated by plotting $\ln(C_t/C_o)$ versus reaction time for the reduction of 4-NP (Figure 6). The slopes of graph indicate pseudo first-order kinetics.⁵ The apparent rate constants for Cu, Pd and Pd@Cu bimetallic nanoparticles were determined from the slopes of linear curve, which were 0.3322 min^{-1} , 0.2689 min^{-1} and 1.812 min^{-1} , respectively. It was observed that the correlation coefficients (R^2) were above 0.99 for all the three types of synthesized nanoparticles. The reduction rate was found to increase 6 times with doping Pd nanoparticles over Cu nanoparticles (Pd@Cu) as compared to monometallic Cu and Pd nanoparticles. The catalyst was recycled 7 times consecutively by centrifugation at 14,000 rpm and

used reduction of 4-NP, with the observation of 1% deficiencies in reuse activity.

Comparative antibacterial properties of Cu, Pd and their bimetallic Pd@Cu nanoparticles

The researchers have been forced to discover some new antibacterial agents to eliminate the risk of multidrug-resistant bacterial strains. For this purpose, we have compared the antibacterial activities of citrate-stabilized Cu, Pd and Pd@Cu bimetallic nanoparticles. In the reported study, different nanoparticles against the eight different strains of bacteria were tested. Among these series of setup, *B. thuringiensis* and *S. aureus* were found to be the positive strains and the rest, *Proteus mirabilis*, *Shigella flexneri*, *Klebsiella pneumonia*, *Pseudomonas aeruginosa*, *S. typhimurium* and *Escherichia coli*, were recorded as the negative strains. With reference to

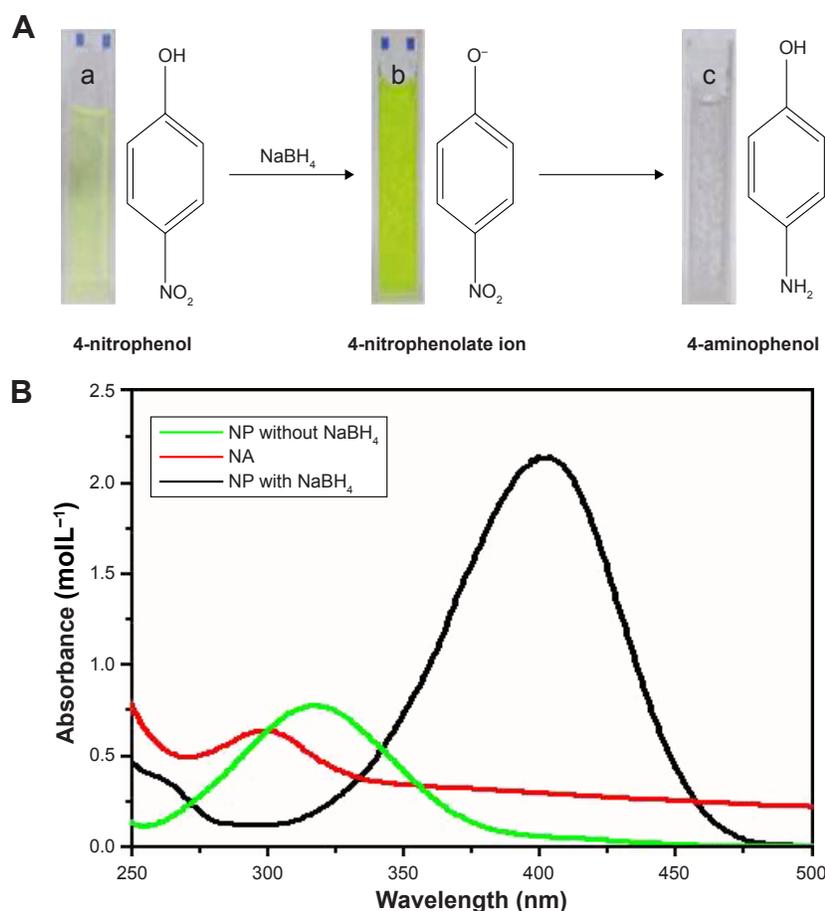


Figure 4 (A) Color change during catalytic reduction of 4-nitrophenol to 4-aminophenol and **(B)** ultraviolet–visible spectra of p-nitrophenol, p-nitrophenolate ion and 4-aminophenol. Reaction conditions: H₂O (3 mL), Cu@Pd (15 μ L, 0.1 mg mL⁻¹), 4-nitrophenol (55 μ L, 0.1 M) and NaBH₄ (150 μ L, 0.1 M). (a) 4-nitrophenol, (b) 4-nitrophenolate ion, (c) 4-aminophenol.

Abbreviations: NA, nitroaminophenol; NP, nitrophenol.

our literature survey, it was found that Cu and Cu-containing nanocomposites had a better antibacterial activity. Therefore, in this work, we have checked out in detail the antibacterial activities of Cu, Pd and their composites through disk diffusion method against the various bacterial strains, that is, Gram^{+ve} and Gram^{-ve}.

According to the diffusion method, the antimicrobial properties of Cu, Pd and Pd@Cu bimetallic nanoparticles were checked against the Gram^{+ve} bacteria *B. thuringiensis* and *S. aureus*. At 37°C after 24 h of incubation, the diameters of inhibition zones were calculated and found that both the Cu and Cu@Pd bimetallic nanoparticles had a better sensitivity over both strains of bacteria except Pd, which has no zone of inhibition in the presence of positive control ofloxacin. Kora and Rastogi have investigated that Pd nanoparticles have no antibacterial activities.³⁷ The zones of inhibition of Cu nanoparticles against *B. thuringiensis* and *S. aureus* were found to be at 14 \pm 0 and 16 \pm 0 mm, respectively, as shown in Figure S2A and B. The positive control plate loaded with ofloxacin has a zone of inhibition at 16 mm. The zone of inhibition of Pd@Cu bimetallic nanoparticles was observed

at 17 \pm 0 and 20 \pm 0 mm. The positive control ofloxacin had presented the zone of inhibitions at about 19 \pm 0 mm.

Similarly, the same nanoparticles were used against Gram^{-ve} bacterial strains *Proteus mirabilis*, *Shigella flexneri*, *Klebsiella pneumonia*, *Pseudomonas aeruginosa*, *S. typhimurium* and *E. coli* to check out in detail the antibacterial properties. The inhibition zones of Cu nanoparticles against the above-mentioned negative bacterial species strains were recorded as 17 \pm 0, 12 \pm 0, 12 \pm 0, 15 \pm 0, 12 \pm 0 and 19 \pm 0 mm, respectively. The positive control sides of ofloxacin inhibition zone against the proposed negative bacterial strains were found as 18 \pm 0, 16 \pm 0, 14 \pm 0, 17 \pm 0, 15 \pm 0 and 20 \pm 0 mm, respectively, as shown in Figure S2C–H.

The inhibition zones of Pd@Cu bimetallic nanoparticles against the six negative bacterial species strains presented above are at 19 \pm 0, 18 \pm 0, 17 \pm 0, 18 \pm 0 and 16 \pm 0, 25 mm, respectively, as shown in Figure S2A–H. The diameter of inhibition zones was different for different bacterial strains for Cu, Pd and Cu@Pd bimetallic nanoparticles. This difference in the diameter of inhibition zone was due to the composition of bacterial cell wall. The zone of inhibition of

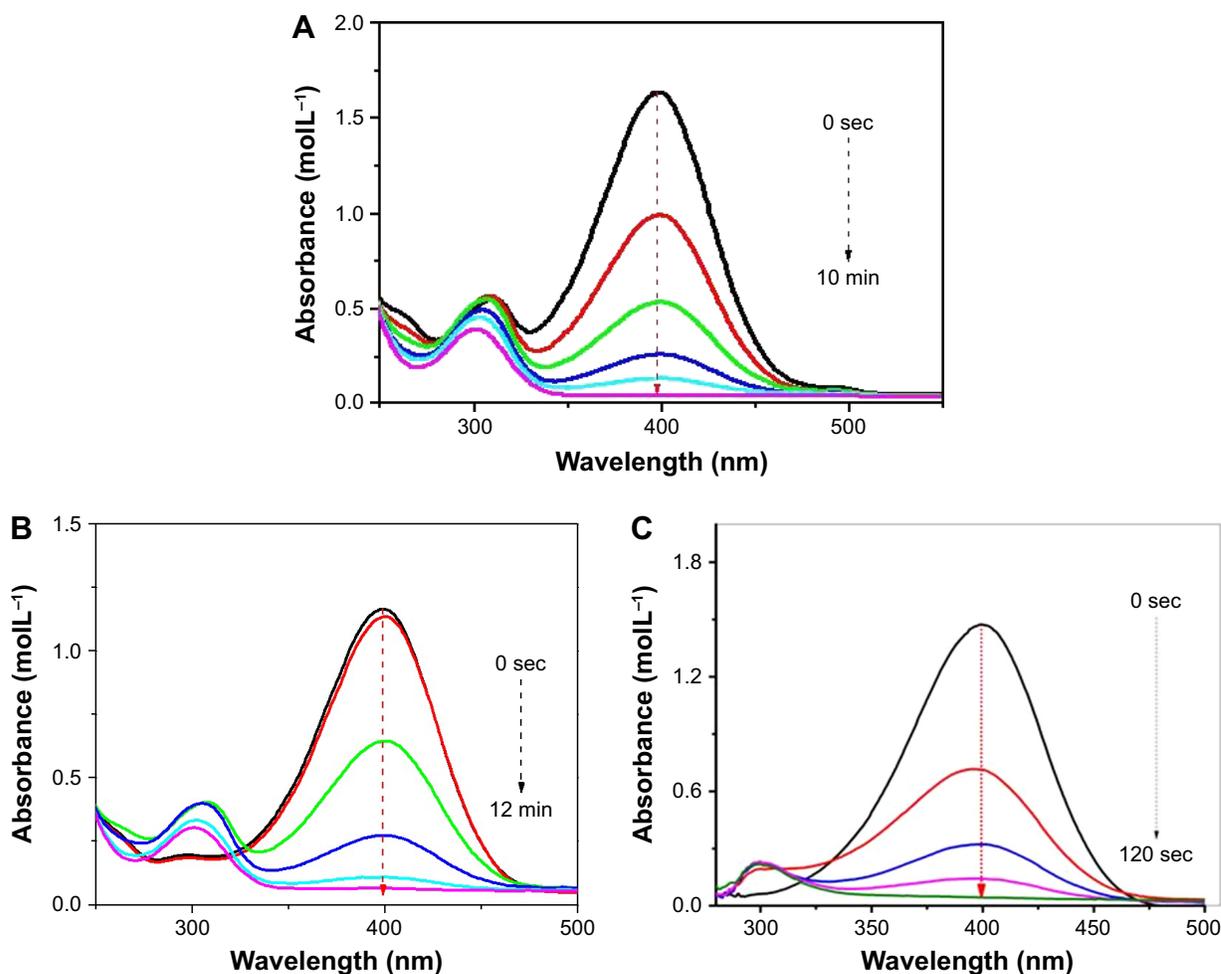


Figure 5 Ultraviolet-visible spectra for the reduction of 4-nitrophenol catalyzed by (A) copper, (B) palladium and (C) palladium@copper bimetallic nanoparticles.

Notes: Reaction conditions: H₂O (3 mL), Cu (15 μ L, 0.1 mg mL⁻¹), 4-nitrophenol (40 μ L, 0.1 M) and NaBH₄ (150 μ L, 0.1 M). (B) Conditions: H₂O (3 mL), Pd (15 μ L, 0.1 mg mL⁻¹), 4-nitrophenol (40 μ L, 0.1 M) and NaBH₄ (150 μ L, 0.1 M). (C) Conditions: H₂O (3 mL), Pd@Cu (15 μ L, 0.1 mg mL⁻¹), 4-nitrophenol (40 μ L, 0.1 M) and NaBH₄ (150 μ L, 0.1 M).

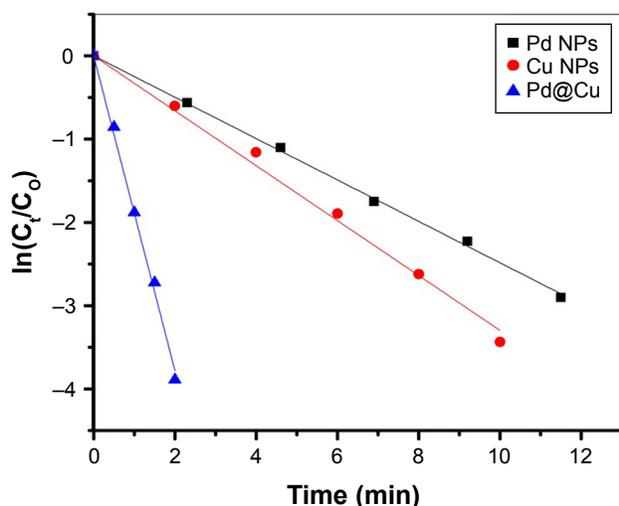


Figure 6 Determination of the reduction rates of 4-nitrophenol as a function of reaction time by using copper, palladium and palladium@copper bimetallic nanoparticles as a catalyst.

Notes: Reaction conditions: H₂O (3 mL), Cu (15 μ L, 0.1 mg mL⁻¹), 4-nitrophenol (40 μ L, 0.1 M) and NaBH₄ (150 μ L, 0.1 M). (b) Reaction conditions: H₂O (3 mL), Pd (15 μ L, 0.1 mg mL⁻¹), 4-nitrophenol (40 μ L, 0.1 M) and NaBH₄ (150 μ L, 0.1 M). (c) Reaction conditions: H₂O (3 mL), Pd@Cu (15 μ L, 0.1 mg mL⁻¹), 4-nitrophenol (40 μ L, 0.1 M) and NaBH₄ (150 μ L, 0.1 M).

The bimetallic Cu@Pd nanoparticles zone of inhibition was reported with maximum range rather than the standard ofloxacin, which shows that these bimetallic nanoparticles are more effective against all the bacterial strains. The citrate-stabilized Pd@Cu nanoparticles have shown maximum antibacterial activities than the reported work in our existing literature.^{38,39}

Conclusion

The present research work focuses its efforts to synthesize the monodispersed Cu, Pd and their bimetallic Pd@Cu nanoparticles via known chemical reduction process in the presence of trisodium citrate as stabilizer. The synthesized monometallic and bimetallic nanoparticles had a remarkable stability and increase in amount with increase in volume of trisodium citrate. The highly catalytic and antibacterial activities were observed with bimetallic Pd@Cu nanoparticles. The catalytic activities of monometallic and bimetallic nanoparticles were gaged on 4-NP and an increase up to 6 times on doping Pd nanoparticles over Cu nanoparticles (Pd@Cu) as

compared to their monometallic nanoparticles was observed. Meanwhile, the antibacterial activities were carried out on two Gram⁺ and six Gram⁻ bacterial strains. The bimetallic nanoparticles had shown the maximum antibacterial activities as compared to monometallic Cu and Pd nanoparticles. Moreover, the antibacterial activities of bimetallic Pd@Cu nanoparticles were found at a maximum level than the used standard antibacterial drug ofloxacin. It is recommended that the developed findings in terms of catalytic property and antibacterial activities of the proposed compounds will enable the use of bimetallic nanoparticles in the field of catalysis and nanomedicine.

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Disclosure

The authors report no conflicts of interest in this work.

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Supplementary materials

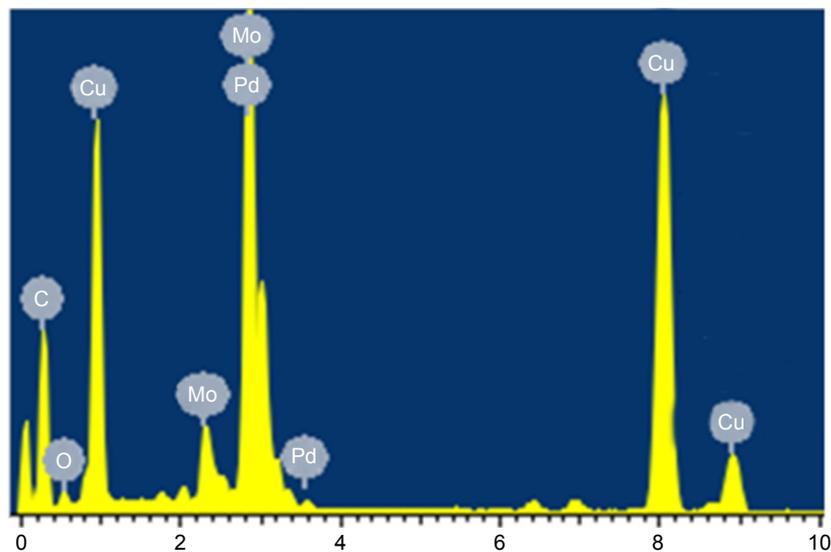


Figure S1 Energy-dispersive X-ray spectrum of Pd@Cu nanoparticles.

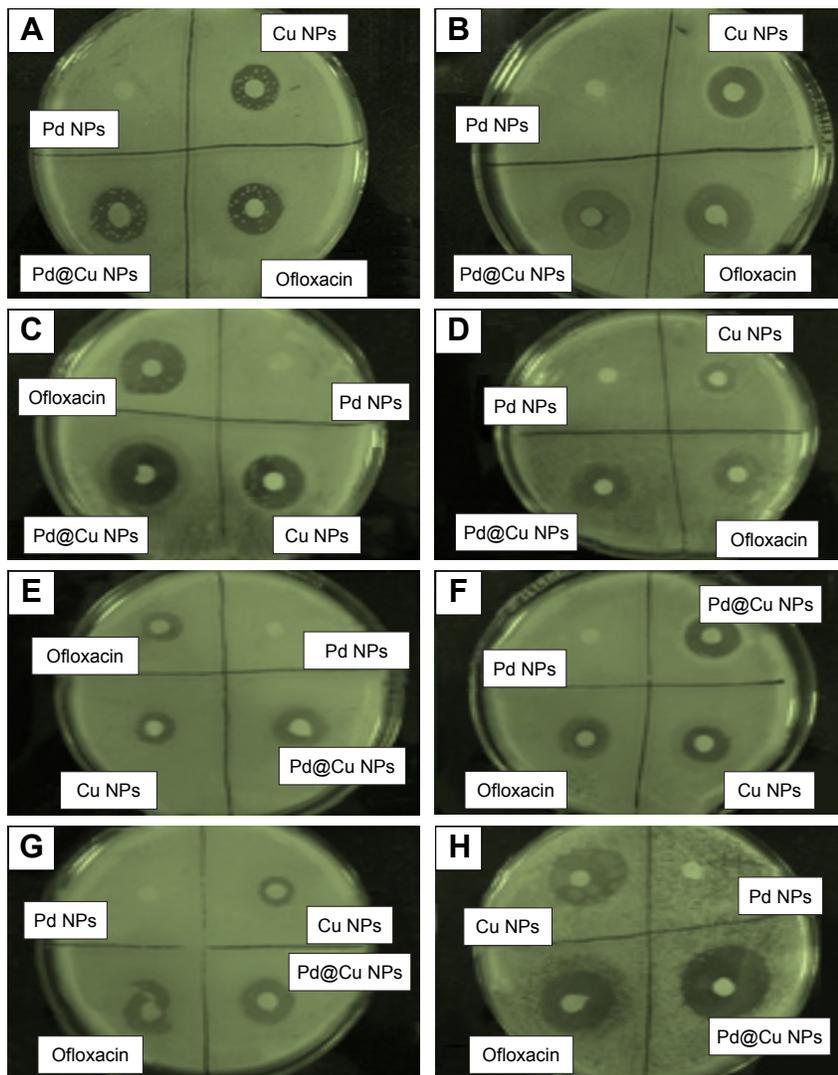


Figure S2 Comparative antibacterial properties of copper, palladium and their bimetallic palladium@copper nanoparticles: (A, B) Gram⁺ve; (C-H) Gram⁻ve bacteria.

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