Biocompatibility and antibacterial activity of nitrogen-doped titanium dioxide nanoparticles for use in dental resin formulations

Abstract: The addition of antibacterial functionality to dental resins presents an opportunity to extend their useful lifetime by reducing secondary caries caused by bacterial recolonization. In this study, the potential efficacy of nitrogen-doped titanium dioxide nanoparticles for this purpose was determined. Nitrogen doping was carried out to extend the ultraviolet absorbance to longer wavelength blue light for increased biocompatibility. Titanium dioxide nanoparticles (approximately 20–30 nm) were synthesized with and without nitrogen doping using a sol–gel method. Ultraviolet–Visible spectroscopy indicated a band of trap states, with increasing blue light absorbance as the concentration of the nitrogen dopant increased. Electron paramagnetic resonance measurements indicated the formation of superoxide and hydroxyl radicals upon particle exposure to visible light and oxygen. The particles were significantly toxic to Escherichia coli in a dose-dependent manner after a 1-hour exposure to a blue light source (480 nm). Intracellular reactive oxygen species assay demonstrated that the particles caused a stress response in human gingival epithelial cells when exposed to 1 hour of blue light, though this did not result in detectable release of cytokines. No decrease in cell viability was observed by water-soluble tetrazolium dye assay. The results show that nitrogen-doped titanium dioxide nanoparticles have antibacterial activity when exposed to blue light, and are biocompatible at these concentrations.

Keywords: titanium dioxide, antibacterial activity, nitrogen doping

Introduction

Despite the long-term environmental and human health concerns related primarily to mercury, amalgam dental restorations are still widely used. This is due to their enhanced durability over composite resins, particularly in the case of large restorations placed in molars. Failure of composite resin restorations is typically the result of two major factors: errors during placement, such as inadequate or uneven curing, and secondary caries caused by shrinkage and breakage of the material leading to colonization of biofilm-forming bacteria in the resin–dentin interface. A 2003 review found that these failures are due to improper placement technique or incorrect choice of materials when they occur within the first 5 years after placement. Failures occurring between 5 and 17 years are typically the result of secondary caries. Overall failure rates during this time period are as high as 45%.

The acidic conditions associated with secondary caries primarily affect the remaining tooth, but not the restoration material. Streptococcus mutans biofilms grown on a resin–dentin interface cause a significant reduction in fatigue resistance of the joint material due to degradation of dentin at the resin–dentin interface. The incorporation of...
of an antibacterial agent into this region, which is inaccessible during typical oral prophylaxis, will reduce dental restoration failure and increase their effective lifetime. Traditional organic molecules used as antibacterial agents are not ideal for this application, as they have limited stability and lifetimes. Inorganic materials with antibacterial activity have generally longer usable lifetimes, and thus have greater potential for successful use in this case.

The development of nanomaterials, materials smaller than 100 nm in at least one dimension, has resulted in advancements in many fields, including the improvement of antibacterial and “self-cleaning” properties of materials such as concrete, glass, and medical devices and materials. TiO$_2$ has been extensively used in the development of these materials due to photocatalytic activity stimulated by ultraviolet (UV) light exposure. Upon excitation with UV light, TiO$_2$ generates reactive oxygen species (ROS) in the presence of water. The ROS generated in this reaction lead to the decomposition of organic compounds and have demonstrated antibacterial activity toward several bacterial strains including Escherichia coli, Pseudomonas aeruginosa, and Staphylococcus aureus.

Nanoparticles, primarily silica and zirconium oxide, are currently used in dental composite resins to increase the strength properties while reducing material shrinkage during curing. There is a clear opportunity to replace these with functionalized nanomaterials that could impart additional benefits, such as antibacterial TiO$_2$. Recently, an antibacterial composite was developed that also had additional benefits, such as antibacterial TiO$_2$. Recently, an antibacterial composite was developed that also had self-healing and remineralization properties using calcium phosphate nanoparticles. Another self-healing composite material was developed using breakable capsules of unreacted monomers, which has shown promising healing characteristics.

Recent research has shown that the light absorption of TiO$_2$ can be shifted to the visible range by doping with a variety of atoms, including nitrogen, fluorine, copper, and silver, resulting in a new class of visible light active materials. This is of particular interest in the development of novel dental materials for two reasons: exposure of gum tissue to blue light is preferable to cytotoxic UV light and most of the dental offices are already equipped with blue light sources used to cure photoactive composite dental resins. The incorporation of blue light–sensitive TiO$_2$ into a dental material could allow for long-term and repeatable release of antibacterial ROS. This is currently not possible with traditional commercially available organic antibiotics, which are consumed upon use.

Techniques for doping of TiO$_2$ to increase the UV and visible light sensitivity initially utilized metals. This was generally unsuccessful due to issues with reproducibility and lack of photocatalytic activity at visible wavelengths, attributed to high carrier recombination rates. Later, nonmetal-doped TiO$_2$ was developed, primarily with nitrogen, which is to date the most successful and widely utilized dopant for the visible light sensitization of titania. Nitrogen-doped TiO$_2$ with codopants such as silver has also been developed with promising characteristics. Metal dopants or codopants are, however, not ideal for eventual clinical use due to the potential for metal leaching and toxicity to the surrounding tissue when used at concentrations necessary for killing pathogenic bacteria.

In this report, nitrogen-doped TiO$_2$ particles (N-TiO$_2$ NPs) were synthesized based on a representative literature source for potential incorporation into a novel dental resin material with light-activated antibacterial properties. N-TiO$_2$ NPs were extensively characterized by various techniques. To determine if the particles are appropriate for further study toward this application, their toxicity under blue light (485 nm) exposure to a model bacterium, E. coli, and their biocompatibility with human gingival epithelial cells (hGEPs) were examined.

**Methods**

**Chemical synthesis and characterization**

**N-TiO$_2$ synthesis**

Synthesis of N-TiO$_2$ NPs was adapted from a recent literature synthesis, wherein a sol–gel synthesis of TiO$_2$ was modified to add ethylmethylamine as a nitrogen source.

Titanium isopropoxide, anhydrous ethanol, hydrogen peroxide (30%, w/v), and ethylmethylamine were all purchased from Sigma Aldrich and used without further purification. The water used in the procedure was purified by a Millipore Milli-Q Advantage A-10 Water Purification System.

Titanium tetraisopropoxide (5 g) was first added to 10 mL anhydrous ethanol in a dry inert (nitrogen) environment. This mixture was added to 50 mL of water while stirring (600 rpm), which forms a white TiO$_2$ precipitate instantaneously. Stirring was continued for an additional 10 minutes after which point the precipitate was washed with water three times by centrifugation (5,000×g).

In a fume hood, 10 mL of 30% (w/v) hydrogen peroxide was added to the washed precipitate while stirring (600 rpm), which results in the formation of an orange transparent solution after several minutes. After dissolution, 0.104 g ethylmethylamine was added and the solution was then stirred...
overnight until it thickened into a gel. The gel was dried for 24 hours at 100°C, and then ground into a powder using a mortar and pestle. The resulting powder was then calcined at 300°C in a muffle furnace for 24 hours. To determine the impact of nitrogen precursor concentration on doping levels, the amount of ethylmethylamine added was varied as 0, 0.104, 0.130, and 0.156 g. The final samples were labeled 1, 1.25, and 1.5 N-TiO₂, respectively, signifying the relative amount of nitrogen added based on the literature synthesis.¹⁸

Importantly, care should be taken, particularly at higher concentrations of ethylmethylamine or if attempting to scale up this synthesis, as ethylmethylamine is unstable and ignitable. While the 1.5 N-TiO₂ NP sample was found to absorb the most visible light, a subsequent batch of these particles ignited during the calcination step, presumably due to an excess of unreacted ethylmethylamine. Also, 1.25 N-TiO₂ NPs were synthesized without any untoward incident, and the majority of experiments detailed below were carried out with this nitrogen loading ratio. When scaling up the synthesis, this issue could potentially occur at lower ethylmethylamine concentrations without proper ventilation.

Particle characterization

Diffuse reflectance UV–Visible (UV–Vis) absorption spectra of all samples were recorded from 400 to 800 nm using a Shimadzu UV-2501PC spectrophotometer with ISR-2200 diffuse reflectance attachment.

Zeta potential measurements were recorded with a Malvern Zetasizer Nano ZS. The pH titrations were performed with an attached MPT-2 titrator supplied with 1 M hydrochloric acid. Titration samples were preadjusted to a pH between 11 and 12 by adding 1 M sodium hydroxide dropwise, sonicated, and allowed to rest for 30 minutes. Three zeta potential measurements were taken at approximately 0.2 pH unit intervals between pH 11 and 4, with 30 seconds resting time between each measurement. Zeta potential measurements include a minimum of 10 runs in monomodal mode and were analyzed using Smoluchowski model.

High-resolution transmission electron microscopy (TEM) images were obtained using a Tecnai-F20 system. Particles were washed twice by centrifugation and replacement of supernatant with deionized water. The particles were then suspended in a dilute ethanol solution and deposited onto lacey carbon–coated copper grids. Primary particle sizes were evaluated using ImageJ software, measuring at least 50 particles for each sample.

Electron paramagnetic resonance (EPR) spectroscopic measurements were carried out on a Bruker X-band spectrometer with a high-sensitivity resonator at room temperature. General instrument settings were as follows: microwave power, 10 mW; modulation amplitude, 1 G; receiver gain 1×10⁵; time constant, 82 milliseconds; and time sweep, 42 seconds. Using the spin trap, 5,5-dimethyl-1-pyrroline-N-oxide (DMPO), radical adducts were identified from dispersed particles in both aqueous and dimethyl sulfoxide (DMSO) solutions.⁶

In a typical experiment, approximately 13–14 mg of doped 1.25 N-TiO₂ in 100 μL distilled water or DMSO suspension was prepared. Equal volumes (25 μL) of TiO₂ suspension (aqueous or in DMSO) and DMPO solution were thoroughly mixed to give a final DMPO concentration of 100 mM. For DMSO experiments, the suspension was bubbled with air, oxygen, or argon for 5 minutes. The resulting suspensions were transferred to 50 μL capillary tubes and placed inside the sample cavity. The suspensions were irradiated inside the sample cavity using a fiber-optic 150 W quartz halogen visible light source. The hyperfine splitting constants of the spin adducts were determined by simulating the spectra using an automatic fitting program.³¹

Powder X-ray diffraction patterns of doped and undoped TiO₂ NPs were collected on a Rigaku Geigerflex diffractometer using Ni-filtered Cu Kα radiation at 40 kV and 25 mA at 2θ from 20° to 70°.

Biological studies

Cell culture and N-TiO₂ NP exposures

Pooled primary hGEPs were purchased from CellnTec Advanced Cell Systems AG (Stauffacherstr, Switzerland) and propagated in CnT-Prime epithelial culture medium (CellnTec Advanced Cell Systems AG) on 100 mm petri dishes coated with 30 μL capillary tubes and placed inside a high-sensitivity resonator at room temperature. General instrument settings were as follows: microwave power, 10 mW; modulation amplitude, 1 G; receiver gain 1×10⁵; time constant, 82 milliseconds; and time sweep, 42 seconds. Using the spin trap, 5,5-dimethyl-1-pyrroline-N-oxide (DMPO), radical adducts were identified from dispersed particles in both aqueous and dimethyl sulfoxide (DMSO) solutions.⁶

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Temperatures for both light exposure conditions ranged between 30°C and 34°C. Biological experiments were performed with N-TiO$_2$ only as TiO$_2$ NPs did not show absorption in the blue light range.

**Cell viability assays**

Mitochondrial activity in hGEPs, as measured using water-soluble tetrazolium dye (WST-1) (Roche, Indianapolis, IN, USA), was assessed after incubation with no NPs (negative control), N-TiO$_2$ NPs, or 1% Triton X (positive control) following 60-minute light exposure. After removal of N-TiO$_2$ NPs by aspiration and culture media replenishment, WST-1 reagent was added to each well (10 μL) of 96-well plates. The plates were briefly tapped to ensure homogeneous reagent mixture and then incubated at 37°C and 5% CO$_2$ for 2 hours before reading at an absorbance wavelength of 440 nm and a reference wavelength of 600 nm.

**Oxidative stress**

ROS formation in hGEPs exposed to increasing concentrations of N-TiO$_2$ NPs was quantified using 5-(and-6)-carboxy-2′,7′-dichlorodihydrofluorescein diacetate, acetyl ester (CM-H$_2$DCFDA) (Molecular Probes, Eugene, OR, USA). hGEPs exposed to Dulbecco’s phosphate-buffered saline only or H$_2$O$_2$ served as controls. N-TiO$_2$ NPs at all concentrations were included in wells without cells to determine potential induction of spontaneous fluorescence of CM-H$_2$DCFDA. Fluorescence was measured immediately after light exposures using an excitation wavelength of 490 nm and an emission wavelength of 535 nm at 0, 30, and 60 minutes postexposure.

**N-TiO$_2$ NP antibacterial activity against E. coli**

The goal of this research was to test the antibacterial activity of N-TiO$_2$ NPs in suspension before moving onto more complex biofilm models. Although *E. coli* is rarely associated with caries formation, other Gram-negative bacteria such as *Porphyromonas gingivalis* are frequently found in the oral cavity and can cause certain oral diseases, including gingivitis. Therefore, *E. coli* was chosen as a model bacterium to represent Gram-negative bacteria. Bacterial suspensions of *E. coli* were grown overnight in lysogeny broth (LB) media and diluted to an approximate concentration of 1×10$^7$ colony forming units/mL in phosphate-buffered saline (PBS). The suspended bacteria were plated in triplicate with the concentrations of particles ranging from 0 to 200 μg/mL. The bacteria were then exposed to ambient room light for 60 minutes as a control or to 485 nm at a constant power of 12.85 mW/cm$^2$ and a distance of 55 mm inside a custom-made light box. After light exposures, the bacteria were recovered for 5 hours at 37°C, serially diluted in PBS, and plated on LB agar plates for colony count measurements.

For confocal microscopy experiments, *E. coli* were suspended at a concentration of 1×10$^8$ colony forming units/mL in 24-well plates in LB media. Bacteria were then pelleted, aspirated, and resuspended with N-TiO$_2$ NPs suspended in LB media at a concentration of 50, 100, and 200 μg/mL. The plates were exposed to ambient or blue light (as described above) for 60 minutes. Bacteria were then stained with LIVE/DEAD BacLight stain (Life Technologies, Philadelphia, PA, USA) by incubating them with 1:1 ratios of propidium iodide and SYTO 9 in LB media for 30 minutes in the dark before imaging with a Nikon Eclipse TE-2000E confocal microscope. In the confocal images, green staining is indicative of viable bacteria, whereas red staining represents dead bacteria.

**Statistical analyses**

All hGEP and *E. coli* experiments were repeated three independent times, with a minimum of three technical replicates per experiment.

Statistical comparisons were performed using GraphPad Prism 6 (GraphPad Software, Inc., La Jolla, CA, USA). Two-way analysis of variance with Tukey’s multiple comparisons tests were used to determine if significant changes existed between light treatments as well as NP doses. Data were represented as mean ± standard deviation. Data were considered statistically significant if $P<0.05$.

**Results**

**Particle characterization**

To determine the effects of nitrogen doping on the absorption spectrum of TiO$_2$, UV–Vis diffuse reflectance experiments were performed. Figure 1A shows the UV–Vis reflectance spectra of TiO$_2$ with various amounts of nitrogen doping. Visible light absorption (from 400 to 800 nm) was enhanced with increasing nitrogen doping. Figure 1B shows Tauc plots of the UV–Vis absorption data, which are used to determine the bandgap of each sample. The bandgap was 3.11 eV for undoped, 1 N-TiO$_2$, and 1.25 N-TiO$_2$. This shifted to 2.99 eV for 1.5 N-TiO$_2$. Because of the adequate visible absorption by 1.25 N-TiO$_2$, and the problems with ignition of the 1.5 N-TiO$_2$ sample, all further experiments were carried out with 1.25 N-TiO$_2$.

To determine whether the addition of nitrogen affected the final crystal structure of the particles, X-ray diffraction...
patterns of undoped TiO₂ and N-TiO₂ were measured. These are shown in Figure 2 with anatase and rutile TiO₂ peaks labeled as “a” and “r”, respectively. It is clear from the peak assignments that nitrogen doping resulted in the formation of primarily rutile phase TiO₂, while undoped TiO₂ was primarily anatase.

Particle size was determined by TEM. The TEM images shown in Figure 3 indicate irregular particle sizes and strong aggregation, which did not break apart after extensive sonication. Both the undoped (Figure 3A and B) and 1.25 N-TiO₂ NP samples (Figure 3C) consisted of aggregates ranging from 100 to 300 nm, made up of smaller primary particles: 33.1±10.3 nm for undoped TiO₂ and 21.8±9.6 nm for 1.25 N-TiO₂.

The influence of nitrogen doping on the surface property of TiO₂ was explored with zeta potential versus pH titrations. The isoelectric point (IEP), or the pH at which the particles have a zero zeta potential, was determined to be at pH 5.2 for undoped TiO₂ NPs, while the N-TiO₂ NPs had the IEP shifted to a lower pH of 4.

EPR spectra were measured to specifically identify the radical species generated from particles upon exposure to visible light (approximately 400–675 nm). Figure 4 shows the hydroxyl adducts formed from 1.25 N-TiO₂ in the presence of DMPO under light irradiation. The concentration of N-TiO₂ was 136 mg/mL. No EPR signal was observed without irradiation (Figure 4A), while the signal formation was evident postirradiation (Figure 4B), which matched the simulation data for hydroxyl radical (Figure 4C).

Superoxide adduct has been shown to persist longer in DMSO than in aqueous solution. Therefore, an experiment was carried out in DMSO to investigate O₂⁻ generation. Also, in order to determine the potential role of dissolved O₂ in O₂⁻ generation, the experiments were performed with air, Ar, or O₂ bubbling throughout the suspension.

Figure 5 shows the EPR using N-TiO₂ NPs in the presence of DMSO. No EPR signal was observed before irradiation, as shown in Figure 5A. The spectra for both air- and O₂-bubbled suspensions in the presence of irradiation (Figure 5B) are characterized by the presence of carbon-centered radical as
determined by a fluorescence ROS assay. Figure 6 illustrates the ROS production of hGEPs after exposure to N-TiO₂ NPs under ambient light (Figure 6A) and 485 nm blue light (Figure 6B). The cells produced ROS in response to blue light in the absence of N-TiO₂ NPs. This was accounted for by normalizing NP-exposed cells to the unexposed controls for both time points and light treatments. ROS generation significantly increased in a dose-dependent manner in the presence of blue light treatment.

To investigate if N-TiO₂ NP and light treatment caused cytotoxicity, as indicated by decreases in cellular metabolism in hGEPs, WST-1 assays, which measure the mitochondrial activity, were performed (Figure 7). Cells were treated with increasing concentrations of N-TiO₂ NPs and immediately exposed to either blue light or ambient light for 60 minutes. Following light exposure, the NPs were aspirated and cellular metabolism was then quantified at 3 (Figure 7A) and 24 hours (Figure 7B) postexposure. After 3 hours, the hGEPs exposed to N-TiO₂ and blue light exhibited significant increases in cellular metabolism when compared to the hGEPs exposed to blue light only. Treatment with N-TiO₂ NPs and ambient light resulted in a slight, yet nonsignificant increase in proliferation as well. Cellular metabolism was significantly elevated in cells exposed to 50 and 100 μg/mL N-TiO₂ NPs for 24 hours, irrespective of light treatment. It appears that N-TiO₂ NPs stimulated cellular metabolism at low and intermediate concentrations, but were not cytotoxic at any concentration tested.

To investigate the potential antibacterial effect of N-TiO₂ NPs and light treatments, the colony counts of E. coli after a 60-minute exposure to ambient or blue light in the presence or absence of N-TiO₂ NPs were quantified (Figure 8). Treatment with blue light alone resulted in a significant decrease in E. coli colonies. Also, a significant dose-dependent decrease

*CH₃ and O₂ (as OOH adducts), as shown in the simulated spectra in Figure 5C.

**Biocompatibility and antibacterial studies**

To investigate the biocompatibility of N-TiO₂ NPs with hGEPs, the oxidative stress response of hGEPs after exposure to N-TiO₂ NPs under ambient light or 485 nm blue light was

**Figure 3** TEM images of (A, B) undoped TiO₂ and (C) 1.25 N-TiO₂.

**Abbreviation:** TEM, transmission electron microscopy.

**Figure 4** DMPO spin-trapping EPR spectra of 1.25 N-TiO₂ aqueous solutions under (A) dark conditions, (B) visible light irradiation, and (C) DMPO simulation (DMPO-OH: \( a_N = a_\beta = -14.9 \) G lit: (DMPO-OH: \( a_N = a_\beta = -14.9 \) G)).

**Abbreviations:** DMPO, 5,5-dimethyl-1-pyrroline-N-oxide; EPR, electron paramagnetic resonance.
Biocompatibility and antibacterial activity of N-doped TiO$_2$ NPs

Effect of nitrogen doping on TiO$_2$ structure

Consistent with literature reports,$^{18-21}$ nitrogen doping of TiO$_2$ resulted in absorption extending well into the visible range of wavelengths. The extent of absorption in the visible range of wavelengths was concentration dependent. Adding more nitrogen precursor (ethylmethylamine) resulted in increased visible light absorbance (Figure 1).

There are conflicting literature reports on the mechanism of blue light absorption of N-doped titania. Shifts in the bandgap of TiO$_2$ have been reported upon nitrogen doping: from 3.45 to 2.75 eV$^{18}$ and from 3.2 to 2.4 eV$^{34}$ Other reports in bacterial survival was noted when $E.\ coli$ were exposed to blue light and N-TiO$_2$ NPs, when compared to exposure under ambient light conditions.

To confirm the antibacterial effect of N-TiO$_2$ NPs, an additional set of N-TiO$_2$ NPs and light exposures were performed, and $E.\ coli$ were then subjected to live/dead staining. These confocal fluorescence microscopy studies demonstrated changes in bacterial viability in response to NP exposure. The images shown in Figure 9 reveal a general increase in dead bacteria (red) compared to live bacteria (green) with increasing particle concentration. Damaged bacteria appear to swell and undergo changes in morphology in response to light and particle exposure.

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indicate that the bulk material bandgap does not significantly shift, and that the absorption of blue light is instead the result of a variety of states at the valence band edge attributed to titanium-site vacancies and undercoordinated oxygen ions.\textsuperscript{19,21} Generally, nitrogen doping by substitution into the crystal structure results in lower energy states closer to the original bandgap, while doping by chemisorption and interstitial incorporation results in deeper trap states. This has generally been found to be dependent on the synthetic method and the concentration of nitrogen precursor, with lower concentrations resulting in substitution and higher concentrations resulting in chemisorption and interstitial doping.\textsuperscript{35}

The UV–Vis absorption spectra presented in Figure 1 provide evidence that the increased visible light absorption
in this synthesis can be attributed to valence band trap states. Most of the samples showed no shift in the main bandgap of the material when calculated by Tauc plots (3.11 eV), but did show a tail extending into lower energies characteristic of trap state formation. The Tauc plot calculation for 1.5 N-TiO$_2$ (2.99 eV) appears to be skewed lower by the increasing amplitude of the trap state peak relative to the main bandgap.

The TEM images presented in Figure 3 show nanosized primary particles of approximately 20–30 nm agglomerated into larger 200–300 nm aggregate particles. It should be noted that the aggregation state of the particles can impact their toxicity, as they have a larger effective hydrodynamic radius than the initial primary particles. As some aggregation was observed, it will be imperative to control dispersal when considering the biological applications for N-TiO$_2$ NPs as the aggregation state may alter the biocompatibility of the final material.

From zeta potential pH titration data (Figure 10), we find that undoped TiO$_2$ particles had an IEP of pH 5.2. This matches well with literature: anatase TiO$_2$ particles were found to have an IEP between pH 5 and 6, depending on their size. The IEP of 1.25 N-TiO$_2$ shifted to a pH of 4. Literature reports show that the crystal structure of titania does not affect the IEP, meaning this shift was not due to the transformation to rutile phase. Nitrogen doping has been shown to decrease the zeta potentials and IEP of TiO$_2$. We, therefore, conclude that the change in IEP for N-TiO$_2$ in the present study was due to the addition of nitrogen functionality to the surface.

Next we investigated the ability of the synthesized N-TiO$_2$ NPs to generate free radicals, which could potentially be exploited to make them an effective antibacterial agent. The appearance of EPR signals upon light irradiation in the presence of N-TiO$_2$ NPs is attributed to the formation of OH$^\bullet$. The formation of O$_2$•$^-$ was not observed due to the short half-life of DMPO-O$_2$•$^-$ which decomposes to DMPO-OH, or the fast dismutation of O$_2$•$^-$ to form H$_2$O$_2$ and the subsequent metal-catalyzed homolytic cleavage of H$_2$O$_2$ to *OH to form DMPO-OH. Superoxide adduct of DMPO has been shown to be more persistent in DMSO; hence, spin trapping was performed in aqueous DMSO solution. Comparison of the relative levels of radicals formed in air- versus O$_2$-bubbled suspensions shows higher levels of O$_2$•$^-$ adduct in O$_2$-bubbled system (93%) than with air bubbling (83%) while higher *CH$_3$ level was observed in air-bubbling (17%) than in O$_2$ (7%) and that these signals are significantly decreased with Ar-bubbling.

**Biocompatibility and antibacterial activity of N-TiO$_2$ NPs**

Advancements in the materials engineering field have led to the release of hundreds of consumer products that contain nanomaterials. Applications capitalizing on the unique ability of TiO$_2$ NPs to form ROS under UV light irradiation are of particular interest. However, UV light is damaging and repeated exposure has been linked to skin damage and tumors caused directly by UV irradiation and indirectly through oxidative stress mechanisms and uncontrolled ROS formation. Furthermore, while the cytotoxicity and oxidative stress responses after exposure to undoped TiO$_2$ have been extensively studied, the biological effects of N-TiO$_2$ NPs exposure are less understood. Thus, in addition to thoroughly characterizing the properties of N-TiO$_2$ NPs in a cell-free environment, we also sought to elucidate the antibacterial effects of N-TiO$_2$ NPs and determine their initial biocompatibility in a normal human cell culture model.

The ROS assays (Figure 6) demonstrate that N-TiO$_2$ NPs exposed to blue light induce dose-dependent ROS generation in hGEPs compared to PBS-negative controls. There was a significant change in ROS production between ambient and blue light treatments after exposure to the highest dosage (200 μg/mL) of N-TiO$_2$ NPs. Although there are relatively few investigations which report on ROS production of N-TiO$_2$ NPs in mammalian cell models, our ROS results match those described in Li et al.’s report, which compared the cytotoxicity of undoped and N-doped TiO$_2$ NPs in HeLa cells. In the aforementioned studies, irradiation with blue light (400–440 nm) for up to 5 minutes resulted in substantial increases in total ROS as well as superoxide and hydroxyl radical formation at similar concentrations to those used in our study. As demonstrated through EPR (Figures 4 and 5) studies, the primary ROS generated under visible light exposure were hydroxyl and superoxide radicals. However, there

![Figure 10 Zeta potential titration versus pH for undoped TiO$_2$ NPs and N-TiO$_2$ NPs.](image-url)
are conflicting literature reports on the production of ROS by nitrogen-doped TiO$_2$. Some reports agree with our finding that superoxide radicals are produced,\textsuperscript{30,48} while others indicate that carbon-centered radicals are produced.\textsuperscript{19} Nonetheless, the presence of free radicals indicates that cells may be exposed to conditions that promote oxidative stress, which could result in adverse cellular responses including cell death. To begin to understand what, if any, effects the N-TiO$_2$ NPs may have on normal human cells, metabolic activity measurements in hGEPs were performed immediately (3 hours) or after a 24-hour recovery period.

Interferences have been reported for several types of assays used to estimate cytotoxicity, proliferation, and mitochondrial activity. Specifically, there is a known binding activity between TiO$_2$ and lactate dehydrogenase, which is the commonly used cytotoxicity assay.\textsuperscript{49,50} TiO$_2$ also caused noncellular reduction in the 3-(4,5-dimethylthiazol-2-yl)-2,5-diphenyltetrazolium bromide assay.\textsuperscript{51} In the current study, WST-1 was selected as an indicator of metabolic activity assay because cell-free testing using NPs alone and the WST-1 reagent did not result in any artifacts or interferences, which suggests that WST-1 is a suitable indicator of cellular metabolism and, therefore, can be used as an indirect measure of cellular viability. The metabolic activity of hGEPs increased significantly within 3 hours of light and N-TiO$_2$ NP exposure, which suggests that the ROS may induce enhanced metabolism in response to these stimuli, due to the cells actively working to neutralize any oxidative stress associated with this treatment. Interestingly, the same trend was noted for cells exposed to N-TiO$_2$ NPs 24 hours after irradiation with ambient or blue light, except at the highest concentration tested. These data suggest that N-TiO$_2$ NPs may generate ROS at levels that stimulate mitochondrial activity and increased cellular metabolism, but the cells are equipped with enough antioxidant mechanisms to protect them and preserve their viability. These data are in agreement with Behzadnia et al.,\textsuperscript{52} who demonstrated no adverse effects when fabrics treated with N-TiO$_2$ NPs were exposed to normal human dermal fibroblasts. Yet, the findings in the present study are also in contrast with the viability and cellular responses reported in other studies. As described earlier, Li et al reported significant increases in ROS.\textsuperscript{48} However, in their study, they also demonstrate decline in HeLa cell mitochondrial membrane potential and significant increases in intracellular Ca$^{2+}$ levels, both of which are indicative of severe cytotoxicity. Furthermore, in a previous study, the same group demonstrated that N-TiO$_2$ NPs are capable of inducing cytotoxicity and micronuclei formation in three cancer cell lines.\textsuperscript{53} There are several notable differences between our study and the studies of Li. One such difference is the intensity and duration of light used among the studies. The present study used 12.85 mW/cm$^2$ for 60 minutes, and Li irradiated with 12 mW/cm$^2$ for 4 hours or 40 mW/cm$^2$ for up to 5 minutes in his studies. Also, the particles were prepared using ammonia gas as a nitrogen source. Finally, the cell culture models employed by these authors were cancer cell lines, which often have various susceptibilities to different stimuli, compared with the normal cells used in this study. As such, the work presented here indicates that while the N-TiO$_2$ NPs induce ROS, the normal cells may be equipped with adequate protective mechanisms to ensure their survival. Further studies are warranted to investigate potential cellular responses after repeated or chronic exposure to both light and N-TiO$_2$ NPs.

Multiple studies have demonstrated the antibacterial effects of undoped TiO$_2$ NPs alone and under UV irradiation.\textsuperscript{14,54,55} Several groups have demonstrated effective bacterial and fungal reduction using N-TiO$_2$ NPs and films.\textsuperscript{56–58} In the present work, a repeatable, dose-dependent (0–200 μg/mL) toxicity toward E. coli resulted upon blue light treatment (Figure 8). An approximately 1-log reduction in bacterial population was observed with the highest concentration of particles. These findings were in agreement with the morphological changes and the decrease in number of viable bacteria observed by confocal microscopy. The antibacterial activity reported here is not as substantial as reported by Cao et al, where 90% reduction was observed in Lactobacillus acidophilus and Candida albicans.\textsuperscript{58} Their study was performed using a much longer 24-hour exposure to visible light and a thin film surface of nitrogen-doped titania. Furthermore, it is possible that ROS scavenging occurred, minimizing the potency of oxidative stress.

**Conclusion**

The N-TiO$_2$ NPs synthesized by this method produced ROS (hydroxyl and superoxide) capable of significantly reducing bacterial counts while maintaining the viability of human epithelial cells, in the presence of water, oxygen, and visible light. Subsequent experiments will determine if the N-TiO$_2$ NPs are toxic to biofilm-forming bacteria (S. mutans) when incorporated into a dental resin composite material. While blue light exposure alone induced ROS production in hGEPs, this is not considered prohibitive toward the eventual use of these particles in dental applications. A current barrier to the
translation of these particles into a real-life application is the length of light exposure and the depth of light penetration required to achieve an impactful biological effect. One-hour light exposures were required to exhibit antibacterial activity at the concentrations used in solution against planktonic bacteria. In the resin formulation, particle concentrations can be >50% by weight of the material. It is expected that light exposure times required to achieve antibacterial activity will be reduced. Therefore, future work is aimed at reducing the irradiation times and enhancing the antibacterial effects while maintaining a balance of oxidative stress that is toxic to bacteria and minimally affecting mammalian cellular responses.

Acknowledgments

The authors wish to thank Ms Shannon Olsen for her lab managerial assistance and assay refinement efforts and Mr Duane Cox for his expert engineering skills used to construct the light box. The authors also wish to acknowledge Mr Tony Yuan and Mr Matt Morgan for their assistance with processing the confocal images. In addition, the authors thank Dr Amer Tiba for his critical feedback and thoughtful discussion.

Disclosure

The views expressed in this article are those of the authors and do not necessarily reflect the official policy or position of the Department of the Navy, Department of Defense, or the US Government. Title 17 U.S.C. §105 provides that “Copyright protection under this title is not available for any work of the United States Government”. Title 17 U.S.C. §101 defines a US Government work as a work prepared by a military service member or employee of the US Government, as part of that person’s official duties. This work was funded by the Naval Medical Research Center’s Advanced Medical Development Program using work unit number G1405. Amber Nagy is an employee of the US Government and Andrew Zane, Ann Marie Digeorge Foushee, and Kristin Flores are contracted by the US Government. This work was prepared as part of our official duties. Ranfang Zuo was supported by Fundamental Research Funds for the Central Universities (Grant No 2652013037). The Hungarian National Research, Development and Innovation Office (NKFIH) Grant Nr. K119442 is acknowledged for support of Antal Rockenbauer.

The authors report no other conflicts of interest in this work.

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