

Highly biocompatible, nanocrystalline hydroxyapatite synthesized in a solvothermal process driven by high energy density microwave radiation

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Abstract: A microwave, solvothermal synthesis of highly biocompatible hydroxyapatite (HAp) nanopowder was developed. The process was conducted in a microwave radiation field having a high energy density of 5 W/mL and over a time less than 2 minutes. The sample measurements included: powder X-ray diffraction, density, specific surface area, and chemical composition. The morphology and structure were investigated by scanning electron microscopy as well as transmission electron microscopy (TEM). The thermal behavior analysis was conducted using a simultaneous thermal analysis technique coupled with quadruple mass spectrometry. Additionally, Fourier transform infrared spectroscopy tests of heated samples were performed. A degradation test and a biocompatibility study in vitro using human osteoblast cells were also conducted. The developed method enables the synthesis of pure, fully crystalline hexagonal HAp nanopowder with a specific surface area close to 240 m²/g and a Ca/P molar ratio equal to 1.57. TEM measurements showed that this method results in particles with an average grain size below 6 nm. A 28-day degradation test conducted according to the ISO standard indicated a 22% loss of initial weight and a calcium ion concentration at 200 μmol/dm³ in the tris(hydroxymethyl)aminomethane hydrochloride test solution. The cytocompatibility of the obtained material was confirmed in a culture of human bone derived cells, both in an indirect test using the material extract, and in direct contact. A quantitative analysis was based on the 2,3-bis-(2-methoxy-4-nitro-5-sulfophenyl)-2H-tetrazolium-5-carboxanilide. Viability assay as well as on DNA content measurements in the PicoGreen test. Indirect observations were performed at one point in time according to the ISO standard for in vitro cytotoxicity (ie, after 24 hours of cell exposure to the extracts). The direct contact tests were completed at three time points: after 24 hours, on day 7, and on day 14 of a culture in an osteogenic medium. All of the tests revealed good tolerance of cells toward the material; this was also shown by means of live/dead fluorescent staining. Both quantitative results and morphological observations revealed much better cell tolerance toward the obtained HAp compared to commercially available HAp NanoXIM, which was used as a reference material.

Keywords: bone regeneration, bone substitute, microwave, HAp

Introduction

Resorbable, bone-compatible materials are required for construction of tissue-engineered bone products which are used when a significant amount of new bone regeneration is required. These cases include severe bone fractures resulting from traumatic injury, fusion repair, void filling following the removal of bone tumors, and several dental indications. The last few years have seen a rapid increase in the number of cases in which bone defects required replacement, with most causes due

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to cancer and aging.^{1,2} The current clinical gold standard for bone replacement therapy is an autograft, which provides the primary factors for effective bone regeneration: osteoconductivity, osteoinductivity, and osteogenesis. However, autografts impose a significant cost and risk for the patient with the need for double surgery.³

Advances in synthetic bone graft substitutes are playing an increasingly important role in large bone defects and gap healing, as they are potentially able to overcome the disadvantages of current medical practices. The primary purpose of tissue-engineered resorbable implants is to induce the natural bone healing process, which otherwise would not occur over critical-sized defects. A major challenge in the development of resorbable bone substitutes is the possibility of providing a material with both good initial mechanical stability as well as fast osteointegration and resorption after implantation. Most of these requirements are satisfied by scaffolds made of hydroxyl acids, such as polylactic acid or poly(lactico-glycolic acid); however, despite their good mechanical properties, low toxicity, and predictable biodegradation kinetics, these materials do not provide a proper surface for cell attachment and proliferation due to the lack of specific cell-recognizable signals.⁴

Bioactive ceramics as a material for bone implants have appropriate osteoconductivity and biocompatibility due to their chemical and structural similarity to the mineral phase of native bone. In addition, the interactions of osteogenic cells with these ceramics lead to good bone bonding ability and bone regeneration.⁵⁻⁷ The most commonly used ceramic materials are various grafts made of submicron hydroxyapatite (HAp), calcium phosphate, and bio-glasses, which are all in the form of particles, powders, paste, or granules. They are suitable for small bone gaps.

On the other hand, for the healing of large bone gaps, biocompatible material with a high resorption rate and good mechanical properties is required. Calcium deficient HAp may play an important role as a material satisfying these requirements.⁸ Chemical properties such as ion exchange or the degradation rate in solution are strongly dependent on the Ca/P ratio. Natural apatite crystals have a calcium deficiency Ca/P ratio near 1.5, which is close to tricalcium phosphate stoichiometry, but, structurally and chemically, it is similar to stoichiometric HAp with Ca/P equal to 1.67.^{8,9} Nonstoichiometric HAp in comparison to synthetic stoichiometric HAp has a higher solubility and degradation rate, and additionally these parameters increase with a decrease in the HAp Ca/P molar ratio.⁹ The first step towards the development of such material is to produce nanoparticles that will

serve as building blocks for the future bio-nanomaterial for bone regrowth scaffold.

In addition to nonstoichiometry, particle size may also determine the material resorption rate and its bioactivity.^{10,11} The high value of the specific surface area (SSA) to volume ratio may significantly increase the material solubility, which is important for achieving resorbable scaffolds.¹¹ Additionally, a decrease in the material grain size improves the osteoblast adhesion to the implant surface.¹²

Many papers have described the HAp synthesis process.¹³⁻¹⁶ One popular method is solvothermal synthesis that is conducted in mills.¹³ This process enables achievement of HAp with an average grain size close to 20 nm, and a surface as measured by the Brunauer-Emmett-Teller (BET) technique around 140 m²/g. Due to the additional substrates used during this process, usually the obtained HAp will have a high number of extra ions such as sodium cations or carboxyl anions, and typically the Ca/P stoichiometry is close to 1.67. In some papers the authors have presented the possibility of nonstoichiometric HAp synthesis with the use of mills, which allows us to obtain spherical HAp particles with a wide grain size distribution.¹⁴

Microwave energy has also been used a few times as a source of energy in HAp synthesis, although not in closed-pressure systems with very high energy density. Traditional microwave heating allows the synthesis of spherical and needle-like shapes, with the lowest average size of 5 × 15 nm and stoichiometry mostly close to 1.67.¹⁵

Several papers have presented the synthesis of HAp with very large surface areas, even close to 300 m²/g. In these cases the authors did not achieve pure HAp, but synthesized a product mix of HAp with surface active substrates and unreacted organic additions.¹⁶

Various processes have been developed to prepare nano-sized apatite powders.¹³⁻¹⁶ Within these studies there appears to be a lack of simple synthesis methods which would result in a fully crystalline calcium deficient HAp nanopowder with a hexagonal structure, a grain size much below 10 nm, and a surface area greater than 200 m²/g.

One potential solution is to drive the reactions with microwave radiation, as this is able to homogeneously and rapidly heat the entire volume of the vessel almost without a temperature gradient in the heated materials. Microwave radiation is able to provide heat to a reaction solution much more efficiently than conventional reactions.^{17,18} This method has other advantages such as fast reaction, high yield, excellent reproducibility, a narrow particle size distribution, and high purity.^{19,20}

The purpose of this work was to obtain HAP nanopowder that provides a degradation rate and biological properties much greater than the currently available HAP powders. This paper reports on the synthesis of HAP nanopowder using a solvothermal reaction with microwave heating. The novelty of this method lies in the extremely short synthesis time, in the range of 90 seconds. This short time is made possible by the use of a special reactor with a high microwave energy density reaching 5 W/mL in a water solution, thereby ensuring a nonequilibrium nanostructure of HAP. The paper presents the chemical, physical, and biological properties of the obtained material, including the chemical composition, structure and morphology, solubility, and biocompatibility.

Materials and methods

Preparation of HAP nanoparticles

The HAP nanocrystals were obtained by the following method.²¹ The starting materials used for the HAP synthesis were calcium hydroxide ($\text{Ca}[\text{OH}]_2$, pure; CHEMPUR, Piekary Slaskie, Poland) and phosphoric acid (H_3PO_4 , 85% solution, pure for analysis; CHEMPUR); the reactants were used in a molar ratio Ca/P of 1.67. To form a reaction solution, phosphoric acid was added drop-wise to the calcium solution at a rate of 0.01 mL every 3 seconds, and the solution was intensively stirred at room temperature. After 30 minutes of additional stirring, the reaction solution was transferred into a capped Teflon vessel and heated using microwave radiation. The microwave reactor was produced by ERTEC (Wroclaw, Poland) and was operated at 2.45 GHz and at a power density adjusted to 5 W/mL. The time, pressure, and microwave power were computer controlled. The synthesis conditions were 1 MPa pressure over 90 seconds. After the synthesis, the obtained powder was washed by anhydrous ethanol and dried in an air flow at a temperature of 30°C for 16 hours. Herein, this resulting material is referred to as “GoHAP.”

The HAP nanopowder NanoXIM201 (Fluidinova, Moreira da Maia, Portugal) was used as a reference material in the degradation and biocompatibility tests. To the authors' knowledge, this commercially available HAP powder had the lowest available particle size on the market. Its characteristics were: a density of 2.93 g/dm³, an SSA of 120 m²/g, an average particle diameter calculated from SSA of 17 nm, and a Ca/P molar ratio as declared by the producer and confirmed by the authors of 1.66. Figure 1 shows a scanning electron microscopy (SEM) micrograph of NanoXIM201 nanopowder, showing an average grain size around 20 nm aggregated in micrometer

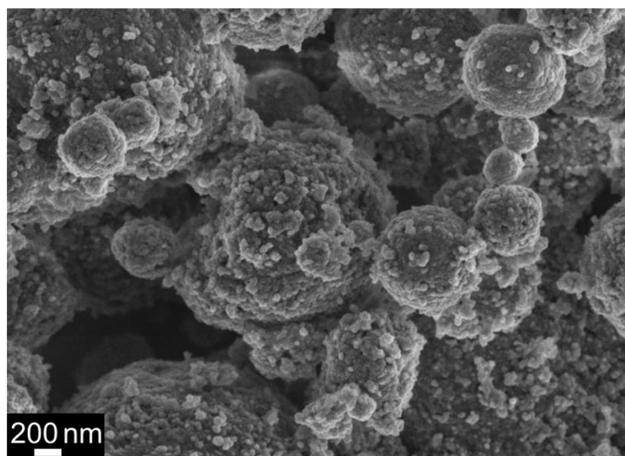


Figure 1 SEM micrograph of NanoXIM201 nanopowder (Fluidinova, Moreira da Maia, Portugal).

Abbreviation: SEM, scanning electron microscopy.

spherical agglomerates. NanoXIM201 is available for medical device applications in the form of nanopowder, which is in compliance with the ISO13779 and ASTM F1185 standards requests for medical application.

Additionally, during the X-ray diffraction (XRD) analysis a plate from a pig's shank bone was used as a reference of natural apatite.

Powder XRD

The phase composition of the reaction products was analyzed by powder XRD (X'Pert PRO diffractometer, Cu K α 1 radiation; PANalytical BV, Almelo, The Netherlands). The patterns were collected at room temperature in the 2 theta range 10°–150° and with a step of 0.03°. Pattern analysis was performed by whole pattern fitting (the Rietveld method) using the DBWS-9807a code.²²

Density and SSA measurements

The density measurements were performed using a helium pycnometer (AccuPyc, model 1330; Micromeritics, Gosford, Australia) using an in-house procedure.²³ The specific surface area of the powders was measured by the BET method (AccuPyc, model Gemini 2360; Micromeritics). The average diameter of the particles was calculated based on the SSA and density, assuming that all of the particles were spherical and identical.²³

Morphology analysis

The morphology of the nanopowder samples was examined with SEM (LEO 1530; Carl Zeiss Meditec AG, Jena, Germany) and transmission electron microscopy (TEM) (JEM2000EX; JEOL, Tokyo, Japan). The TEM

investigations, high-resolution TEM, and selected area electron diffraction were conducted at 200 kV. The specimens for the TEM observations were prepared by dropping the methanol particle dispersion, created by an ultrasonic technique, on a carbon film supported on a 300-mesh copper grid. Additionally, TEM studies were used to determine the nanoparticle size distribution. The grain size histograms were obtained by considering a region of a sample having about 250 nanocrystals and approximating the shape of each nanocrystal by a sphere. The obtained histograms were fitted to either normal or log-normal distributions (Chi-squared test and Pearson's coefficient).

Chemical composition analysis

The chemical composition of the powders was determined by inductively coupled plasma optical emission spectrometry (ICP-OES), with induction in argon plasma (Jobin-Yvon model 138 Ultrace; HORIBA Scientific, Kyoto, Japan).

Degradation test

The determination of material solubility was performed according to norm ISO 10993-14: biological evaluation of medical devices – identification and quantification of degradation products from ceramics. The material was tested in the form of a 6 mm disk (each disk was around 35 mg and was created from the nanopowder using a laboratory hydraulic press under 5 MPa pressure) and was placed in 200 mL of buffer solution (tris(hydroxymethyl)aminomethane hydrochloride [TRIS-HCl] buffer with pH 7.4 ± 0.1). A simulation solution test, lasting 28 days, was conducted without buffer solution exchange during the entire test. All calcium ion concentration measurements were conducted with an ion-selective electrode (Schott ProLab 2000 GLP equipped with calcium ion electrode type CA60; SI Analytics GmbH, Mainz, Germany).

Thermal analysis

The thermal analysis of the HAp samples (approximately 5 mg) was conducted by thermogravimetry and differential scanning calorimetry measurements (TG-DSC) – (STA 449 F1 Jupiter; NETZSCH-Feinmahltechnik GmbH, Selb, Germany). The samples were heated at a rate of $5^\circ\text{C}/\text{minute}$ to 1350°C under a helium gas flow. In the same time the released gases were analyzed by quadrupole mass spectrometry (403 Aëolos; NETZSCH-Feinmahltechnik GmbH).

Additionally, the samples before and after heating were examined by a Fourier transform infrared spectroscopy (FTIR) spectrometer (Bruker Optics Tensor 27; Bruker Corporation, Billerica, MA, USA) equipped with a diamond

attenuated total reflectance (ATR) accessory. The diamond ATR-crystal was soldered onto a ceramic plate, resulting in a chemically inert and mechanically rugged sampling area. All ATR-FTIR spectra were recorded at room temperature in the $4000\text{--}400\text{ cm}^{-1}$ range. The spectral resolution and accuracy were 4 cm^{-1} and $\pm 1\text{ cm}^{-1}$, respectively.

Biocompatibility study in vitro

Isolation and culture of cells used

for biocompatibility study in vitro

Human bone-derived cells (HBDC) were isolated from bone fragments which were amputated from the distal end of the femur during orthopedic surgery and which would have otherwise been discarded. The cells investigated in separate experiments were derived from two different donors. The team obtained the appropriate consent of the Bioethics Committee of the Polish Department of Health to use such biological material for experimental purposes (decision number KB/74/2005) and the donors provided informed consent.

The HBDC isolation procedure was based on the protocols described by Gallagher et al with modifications.^{24,25} After confluence was reached, cells were detached from the culture flasks using collagenase XI S (200 U/mL) (Sigma-Aldrich, St Louis, MO, USA) and 0.25% trypsin suspension, and were used for experiments. Cells were grown in a Dulbecco's modified Eagle's medium ([DMEM] Gibco; Life Technologies, Carlsbad, CA, USA)-based medium supplemented with 10% fetal bovine serum (Gibco), 1% L-glutamine (Gibco), 1% antibiotic-antimycotic mixture (Gibco), and L-ascorbic acid 2-phosphate (Sigma-Aldrich) at a concentration of $34\text{ }\mu\text{g}/\text{mL}$.

Cell culture on the surface of the tested materials

GoHAP and the reference NanoXIM samples were prepared in the form of disks (6 mm diameter, fit to the 96-well culture dishes) using the same disk preparation procedure as described in the degradation test.

Prior to use, the samples were sterilized by irradiation (25 kGy dose) and pre-wetted for 24 hours in a culture medium of the same composition as used in the cell culture. The appropriate samples were placed at the bottom of the 96-well or 24-well tissue culture dishes. The first system was used for quantitative assays while the second one was used for morphological observations. Cells were seeded onto the surface of the examined materials at a density of $2.5 \times 10^4\text{ cells}/\text{cm}^2$. Cells that were cultured on the bottom of tissue-culture polystyrene dishes (TCP) served as a positive control. For the first 24 hours, cells were cultured in a medium

identical with the one used at the isolation stage. Next, the medium was replaced by a commonly used osteogenic medium, ie, it was additionally supplemented with 10 nM dexamethasone (Sigma-Aldrich) and glycerol-2 phosphate (Sigma-Aldrich) at a concentration of 10 mM. All measurements and morphology assays were performed on Day 1 (ie, after the first 24 hours of culture in an osteogenic medium) and then subsequently on Days 7 and 14.

Morphological observations of cells in direct contact with the material

Cell morphology observations in light microscope

The HBDC's morphology was examined both on the surface of the investigated materials (in order to observe cell spreading in direct contact with the surface of the materials), and on the plastic next to the samples. The latter allowed for noninvasive monitoring of the cells' wellness during the entire experiment, as well as observation of the cells that floated from the surface of the material and stayed in the same culture adjacent to the investigated material. The observations were performed by phase contrast microscopy (Eclipse TE2000-U; Nikon Corporation, Tokyo, Japan) connected to a Nikon Digital Sight DS-U1 camera. At least two samples of each type were examined at each time point.

Live/dead viability assay

The viability of the cells cultured on the surface of the investigated samples was qualitatively assessed using live/dead fluorescent staining, based on fluorescein diacetate (FDA) and propidium iodide (PI). Both staining kits were purchased from Invitrogen (Carlsbad, CA, USA), and the staining was performed according to the manufacturer's instructions. The viable cell cytoplasm was labeled green, while the nonviable cell nuclei were labeled red. The principle of the test is that hydrophobic FDA enters a living cell by passing through the hydrophobic cell membrane, and it is then converted to hydrophilic fluorescein by esterase. Fluorescein slowly accumulates in the cell since there is low permeability of the membrane to fluorescein. Accumulation is not possible in dead cells because of the membrane perforation.²⁶ PI intercalates into double-stranded nucleic acids. It is membrane-impermeant and generally excluded by viable cells, but it can penetrate the cell membranes of dying or dead cells. After staining, the cells were covered with the culture medium. Live cells were observed under the microscope in blue light (wavelength of excitation/emission: 494 nm/518 nm) and dead cells were observed in green light (wavelength of excitation/emission: 536 nm/617 nm).

SEM

After viability staining, the osteoblasts grown on the samples were fixed in the 2.5% glutaraldehyde solution in a 0.2 M cacodyl buffer, washed, dehydrated through a graded ethanol series, and air-dried. An SEM examination was performed using SEM microscopy as in the nanopowder morphology studies.

A morphological analysis was performed in two independent experiments using cells from two different donors.

Quantitative assays of the cells cultured on the surface of tested biomaterials

Viability assay

The 2,3-bis-(2-methoxy-4-nitro-5-sulphophenyl)-2H-tetrazolium-5-carboxanilide (XTT) assay was applied. This assay is based on the capacity of mitochondrial dehydrogenase enzymes in living cells to convert the XTT substrate (Sigma-Aldrich), yellow tetrazolium salt (2,3-bis-(2 methoxy-4-nitro-5-sulphophenyl)-5-[(phenylamino)carboxyl]-2H-tetrazolium hydroxide) into a water soluble product, which is orange formazan dye.²⁷ The absorbance of the supernatant (the final product of the reaction) was read at 450 nm in the enzyme-linked immunosorbent assay reader (FLUOstar OPTIMA; BMG LABTECH GmbH, Ortenberg, Germany). The results of the XTT assay are believed to be proportional to the number of living cells.²⁸ The test was performed for cells cultured on the GoHAP and NanoXIM samples, as well as in the TCP control. In most cases, six wells of each type were examined at every time point. The measurements were performed in two independent experiments.

DNA assay

After the XTT assay, the cells were washed twice with phosphate-buffered saline and then treated with 120 μ L of lysis buffer (0.1% Triton X-100; Sigma-Aldrich) in 10 mM Tris-HCl (Sigma-Aldrich), pH 7.4, for 12 hours (overnight) with constant stirring at 4°C. After incubation the plates with cell lysate were stored at -20°C to increase the effectiveness of cell lysis. The PicoGreen assay was used, which is based on the PicoGreen reagent (Invitrogen). The assay uses an asymmetrical cyanine dye, which as a free dye does not fluoresce, but when bound to double-stranded DNA it exhibits a 1000-fold fluorescence enhancement. On the day the measurements were planned, the cell lysate was thawed at 4°C. The PicoGreen assay was performed according to the manufacturer's instructions. The fluorescence was read at a wavelength of excitation/emission of 485 nm/538 nm. The data were related to the DNA standard curve. The test

was performed for cells cultured on both the GoHAP and the NanoXIM samples as well as on the TCP control. In most of the cases, six wells of each type were examined at every time point. The measurements were performed in two independent experiments.

Biocompatibility examination of the extracts from the tested biomaterials

The assays were carried out in accordance with ISO Standard 10993-5:2009.²⁹

The culture medium was chosen as an extraction vehicle. Sterile powders of GoHAP and NanoXIM (irradiation, 25 kGy) were suspended at a concentration of 50 mg/mL. Extract of corundum (Al_2O_3), which is well known as an inert biomaterial, served as a negative control.^{30,31} Extraction was performed for 24 hours with constant stirring at 37°C. After this time, the extracts were centrifuged (MPW-35OR, 10 minutes, 5000 rpm/minute). Supernatants were collected and centrifuged two more times to obtain the clear extracts used in the experiments. The Triton X-100 (Sigma-Aldrich) 0.1% solution in a culture medium served as a positive control.

The HBDCs were seeded at a density of 1.1×10^4 cells per well in 96-well plates, and cultured in DMEM supplemented with 10% fetal bovine serum, 1% antibiotic–antimycotic mixture, 1% L-glutamine, and 34 $\mu\text{g}/\text{mL}$ L-ascorbic acid 2-phosphate. When they reached confluence covering 70% of the well surface, the medium was changed for prepared extracts in a volume of 200 μL per well. Cells were exposed to the following extracts concentrations: 100%, 50%, 25%, and 12.5%. After 24 hours cell morphology was examined under a phase contrast microscope, an XTT assay was performed. Then, cells were lysed and DNA was measured by means of a PicoGreen assay. At least five wells were tested in all groups and the entire procedure was repeated in two separate experiments.

Statistical analysis

All values were presented as mean \pm standard deviation. The differences between experimental groups were statistically assessed using the Kruskal–Wallis Test (nonparametric ANOVA) with Dunn's multiple comparisons test. $P < 0.05$ was considered to indicate statistically significant differences.

Results and discussion Nanoparticles characterization

A typical density for HPA micropowder is $3.05 \text{ g}/\text{cm}^3$.³² In the case of GoHAP the density was lower by almost 4% and it

was $2.93 \text{ g}/\text{cm}^3$. The specific surface area was $236 \pm 5 \text{ m}^2/\text{g}$, and, assuming that the particles were spherical, the average particle diameter calculated on the basis of SSA measurements was 9 nm. The lower density of nanoparticles than in conventional material is in agreement with the literature, which indicated that in the case of nanoparticles, there is a reduction in the material density caused by high contribution of surface layers, which are less densely packed than the bulk.³³

The XRD analysis was conducted on GoHAP, commercial HAP (NanoXIM), and a plate from a pig's shank bone (Figure 2). The XRD results confirmed that the material is a pure well crystalline hexagonal HAP with nano-sized particles. No indication of the presence of other crystalline phases or any amorphous component was found. The lattice parameters were: GoHAP $a = 9.43(4)$, $c = 6.86(5)$; NanoXIM $a = 9.42(3)$, $c = 6.90(5)$; and bone $a = 9.42(5)$, $c = 6.87(7)$. Other structure parameters, except NanoXIM powder, could not be derived with adequate precision, due to the considerable size anisotropy and a very small crystal size. NanoXIM's average particle size calculated by the Scherrer equation was 17 nm. Comparison of the diffraction patterns of the analyzed samples suggested that the crystallites in GoHAP were very similar in size to those in the bone, while those in NanoXIM were considerably larger.

An ICP-OES analysis indicated that the Ca/P molar ratio was 1.57, which was close to the natural bone molar ratio of the calcium and phosphorus elements. Additionally, GoHAP also contains magnesium (0.22 wt%), silicon (0.025 wt%), sodium (0.0125 wt%), and manganese (0.0075 wt%).

The morphology investigation was conducted with SEM and TEM techniques. Figure 3 shows SEM micrographs of the obtained HAP nanoparticles. Strong homogeneity of the obtained nanopowder was observed. The average particle sizes of the synthesized HAP were in the range of 5–20 nm.

A TEM examination showed that the GoHAP has space group P6₃/m of a hexagonal crystal structure with the following parameters: $a = 9.424(4) \text{ \AA}$ and $c = 6.879(4) \text{ \AA}$.³⁴ The analysis did not identify any additional structures in the tested sample. Figure 4 shows the selected area electron diffraction with identified planes. All particles had a needle-like shape (Figure 5), with an aspect ratio (calculated as the ratio of length to width) in the range of 2 to 5. The average grain size for the individual samples was $5.6 \pm 1.1 \text{ nm}$ (Figure 5).

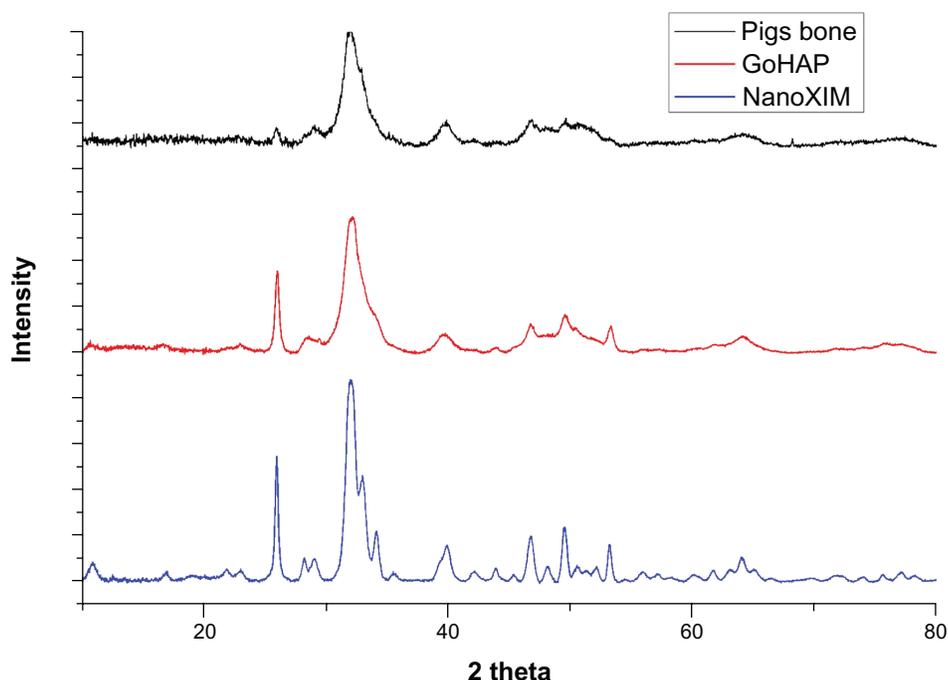


Figure 2 XRD patterns of HAp particles: in pigs bone, GoHAP, and NanoXIM (Fluidinova, Moreira da Maia, Portugal).
Abbreviations: XRD, X-ray diffraction; HAp, hydroxyapatite.

Thermogravimetric analysis

DSC-TG-MS analysis

Figure 6 shows the results of the GoHAP thermal behavior during heating up to 1350°C. The DSC curve shows two endothermic transformations with the onset temperature 375°C (herein called the thermal transformation) and 790°C (herein called the second thermal transformation) (Figure 6A). The loss of weight (Figure 6) from room temperature to the first thermal transformation was 7.4 wt%

and it was caused by the release of water. The first transformation ($\Delta H = -671.2$ J/g) caused 1.9 wt% loss due to the dehydroxylation of material.¹³ At 790°C ($\Delta H = -870.9$ J/g), the second endothermic transformation of HAp took place. After the second thermal transformation, further heating to the 1350°C caused 3.4% weight loss, which was linked to further dehydroxylation. The total amount of released gases from room temperature to the 1350°C was about 15 wt%.

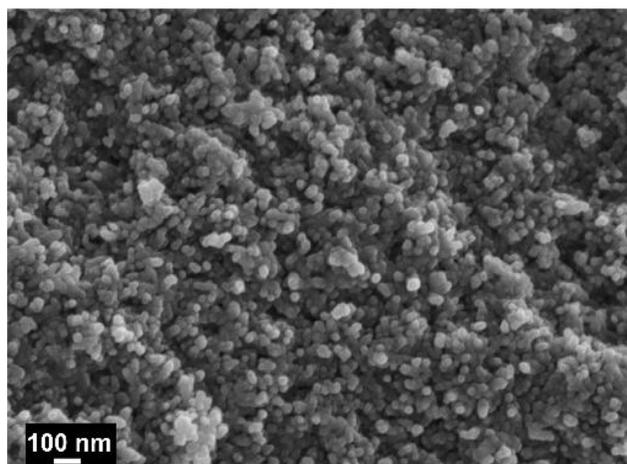


Figure 3 SEM micrograph of GoHAP nanoparticles shows homogeneity of HAp nanopowder.

Note: 100 k zoom (energy 2.00 kV).

Abbreviations: SEM, scanning electron microscopy; HAp, hydroxyapatite.

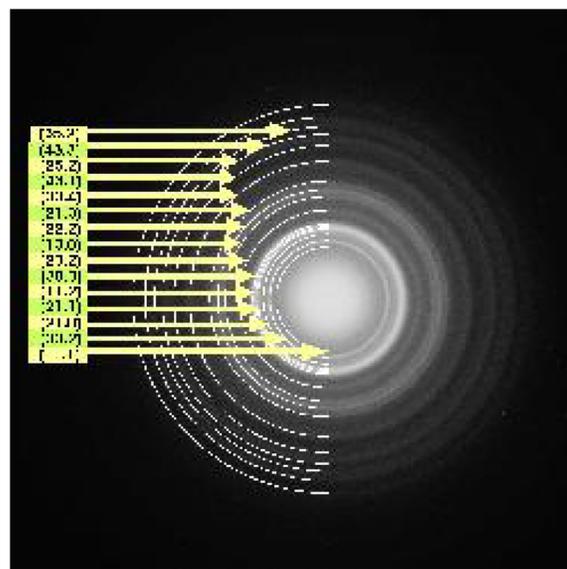


Figure 4 SAED with the identified planes.

Abbreviation: SAED, selected area electron diffraction.

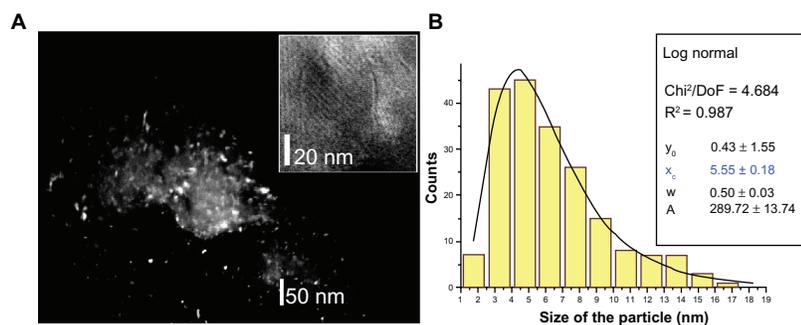


Figure 5 TEM investigation of GoHAP powder. **(A)** Dark field TEM image of GoHAP particles. **(B)** Histogram of the particle size distribution. **Abbreviation:** TEM, transmission electron microscopy.

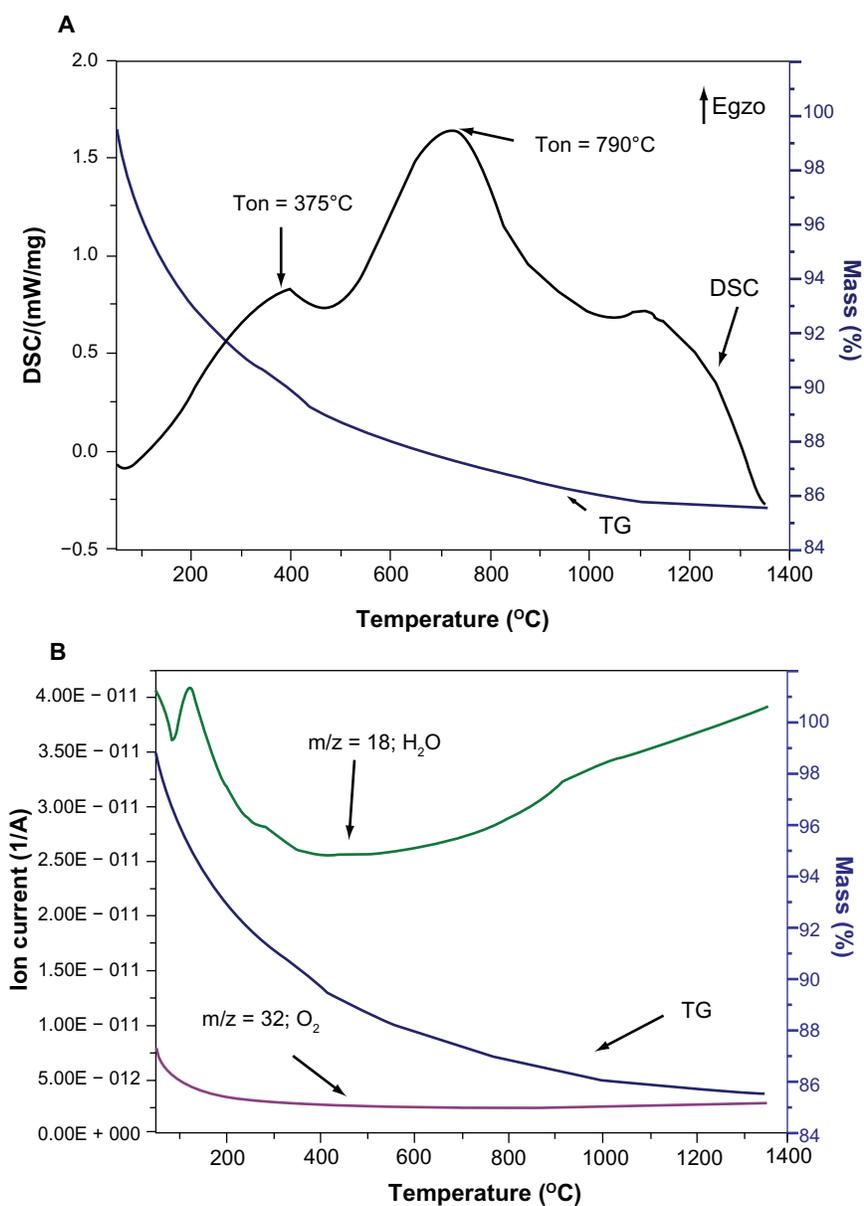


Figure 6 TG analysis of GoHAP powder. **(A)** DSC-TG results of GoHAP heated with a linear ramp 5°C/minute to 1350°C in a helium atmosphere. **(B)** TG-MS results of GoHAP. **Note:** The mass signal of H₂O and O₂ are shown.

Abbreviations: DSC, differential scanning calorimetry; TG, thermogravimetry; MS, mass spectra.

FTIR analysis

The GoHAP samples as-synthesized and samples after heating at a rate of 5°C/minute to different temperatures by the ATR technique of FTIR were investigated. Figure 7A shows the FTIR spectra in the 4000–450 cm⁻¹ range, with the spectral whole (about 2300–2100 cm⁻¹) characteristic of diamond crystal. Additionally, Figure 7B shows the magnification of the spectra for clarity.

The as-synthesized HAP spectra showed typical bands coming from PO₄³⁻. These bands were broadband with a maximum about 1022 cm⁻¹, and another band with two maxima at 601 and 561 cm⁻¹ in the low wavenumbers (Figure 7B). Additionally, bands coming from adsorbed water at 3294 and 1666 cm⁻¹ and from CO₂ coming from air at 2977–2848 cm⁻¹ can be also seen (Figure 7A).^{35,36} Furthermore, a peak at 804 cm⁻¹ origin from the OH⁻ deformation band in HAP (Figure 7B) was observed.

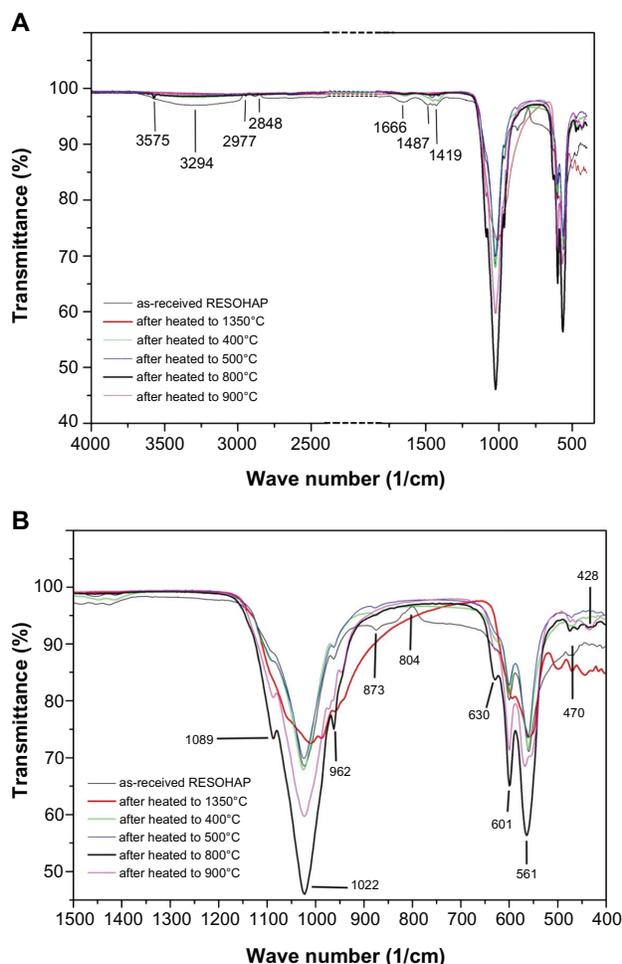


Figure 7 FTIR spectra of the investigated hydroxyapatite samples. Magnification in the wavenumber range 4000 to 400 cm⁻¹ (A) and 1500 to 400 cm⁻¹ (B).

Abbreviation: FTIR, Fourier transform infrared spectroscopy.

After heating the sample above the first thermal transformation temperature (400°C and 500°C), bands were observed coming from adsorbed water (3294 and 1666 cm⁻¹), carbonate (1487–1419, 873 cm⁻¹), and noticeable PO₄³⁻ vibrations that consisted of asymmetric stretching bands P-O ν₃ (1022 and 1089 cm⁻¹), antisymmetric bending bands O-P-O ν₄ (561 cm⁻¹) and stretching bands P-O ν₁ 962 cm⁻¹, symmetric bending bands O-P-O ν₄ 601 cm⁻¹, and bending bands O-P-O ν₂ at 470 cm⁻¹ (Figure 7B).^{13,35}

The spectra of the sample heated above the second thermal transformation (800°C) were characterized by additional bands coming from OH⁻ vibrations, which are typical for stoichiometric HAP.^{37,38} These were the symmetric stretching bands at 3575 cm⁻¹ (Figure 7A) and bending bands at 630 cm⁻¹, which are also characteristic of the sample heated to 900°C. The material after heating to 1350°C presented a fade of the OH⁻ and CO₃²⁻ signals. Only two typical bands of PO₄³⁻ were observed; one broadband at 1022 cm⁻¹, and a second with two clear maxima at the lower wavenumbers 601 and 561 cm.

A thermal stability investigation clearly showed that the obtained material is stable in the temperature range room temperature to 1350°C with two endothermic transformations, the first at a temperature of 375°C and the second at 790°C. Before and during the first transformation, the material loses weight due to water desorption and material dehydroxylation. The TEM investigation did not indicate any changes in the material structure caused by the first transformation. In comparison to the starting material, it was seen that the average grain size increased to 14.4 nm, and that the powder changed its morphology from needle-like grains to a mixture of different shapes, ranging from spherical to needle-like (Figure 8). An ICP-OES analysis of the sample after the first transformation indicated a Ca/P molar ratio after heating of 1.59, which was slightly larger than the ratio in the starting material (1.57). Furthermore, in comparison to the stoichiometric HAP powders with a Ca/P molar ratio equal to 1.67, it showed deep nonstoichiometry. The second transformation was also connected with the further dehydroxylation of the material. A FTIR analysis at temperatures over 800°C indicated the presence of additional bands coming from OH⁻ vibrations, which are typical for stoichiometric HAP.

Degradation test

The tested samples (GoHAP, NanoXIM, GoHAP after the first thermal transformation – herein called GoHAP 375, and a blind sample that was included to measure background changes) were incubated according to norm ISO 10993-14 in

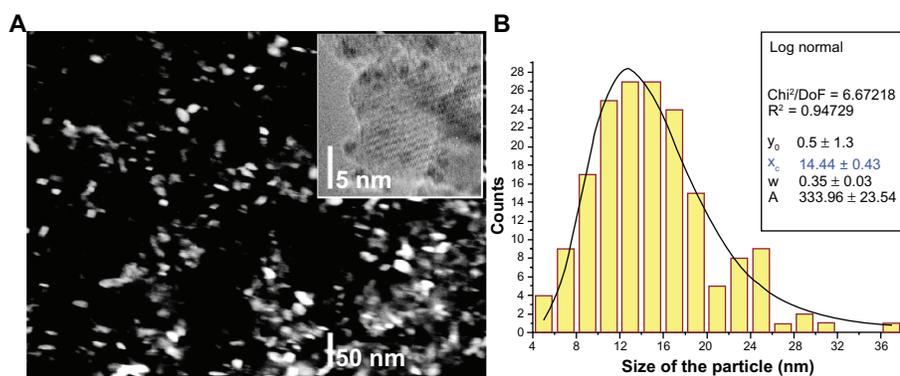


Figure 8 TEM investigation of GoHAP powder after the first thermal transformation. (A) Dark field TEM image of GoHAP after thermal transformation. (B) Histogram of the particle size distribution.

Abbreviation: TEM, transmission electron microscopy.

a TRIS-HCl solution at 37.0°C for 28 days. The initial pH of the test solution was 7.49, the conductivity was 6.68 mS/cm, and the calcium ion concentration was 0.466 $\mu\text{mol}/\text{dm}^3$. Along with the degradation of the samples, the pH, conductivity, and calcium ion concentration values were measured three times per week. The results of the calcium ion concentration changes are shown in Figure 9.

The results showed that in each day GoHAP increased its solubility slightly faster than NanoXIM and GoHAP 375. After this period the calcium ion concentration in the sample with NanoXIM remained almost constant; after 28 days of testing its value was 40.37 $\mu\text{mol}/\text{dm}^3$, which could be recalculated based on weight solubility to 4.0 mg/dm^3 . After 3 weeks of testing GoHAP had twice as much solubility as NanoXIM, and furthermore, in week 4 there was an

acceleration of material solubility, which after 28 days of testing led to the achievement of 199.6 $\mu\text{mol}/\text{dm}^3$ calcium ion concentration, equal to 20.0 mg/dm^3 . The GoHAP 375 sample achieved solubility 4.3 mg/dm^3 after 28 days which is close to NanoXIM results. The conductivity of the solution for GoHAP increased rapidly in the first 3 weeks of the test, to a value of 7.11 mS/cm, and decreased in week 4 to 7.00 mS/cm. The solution with the NanoXIM sample increased its conductivity at the same pace throughout the test to achieve a final value of 7.01 mS/cm, without any rapid changes during the 28 days of experimentation. For the pH measurements there were no crucial differences between the tested materials. The starting pH was 7.49 and its final value ranged from 7.55 for NanoXIM to 7.56 for GoHAP.

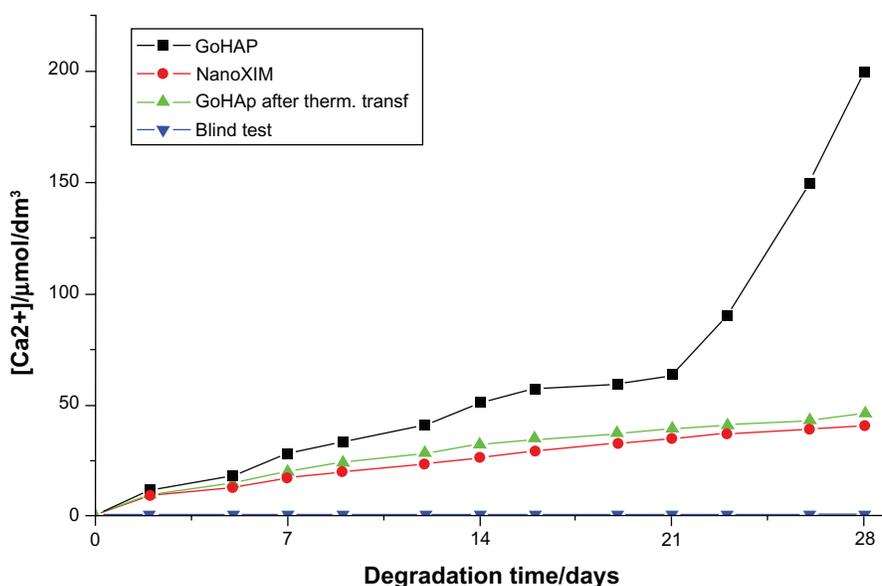


Figure 9 Calcium ion concentration changes for GoHAP, NanoXIM nanopowder (Fluidinova, Moreira da Maia, Portugal), and GoHAP after the first thermal transformation, and a blind test.

Abbreviation: HAp, hydroxyapatite.

The degradation test was additionally followed by a determination of the weight loss after 28 days of testing. During the test, NanoXIM lost 4% of its original weight, resulting in 7.0 mg/dm^3 of solubility, and GoHAP lost over 22% of its initial weight, equal to 38.5 mg/dm^3 . In both the calcium ion concentration and the weight loss measurements, GoHAP had five times greater solubility than NanoXIM. The solubility measured by weight loss was in all cases almost twice as large as the solubility measured by changes in the calcium ion concentration, probably due to nanopowder dispersing. As the SEM investigations showed, the samples of GoHAP after testing were deeply cracked (Figure 10A) but kept their original disk shape. The NanoXIM disks were so weak after the test, that it was impossible to transfer them in one piece to a microscope (Figure 10B shows a fragment of original disk), but their surfaces were not as regularly cracked as the GoHAP disks. GoHAP 375 had a solubility value comparable to NanoXIM (8.1 mg/dm^3), and its disk kept its shape and had a low cracking level after the test period.

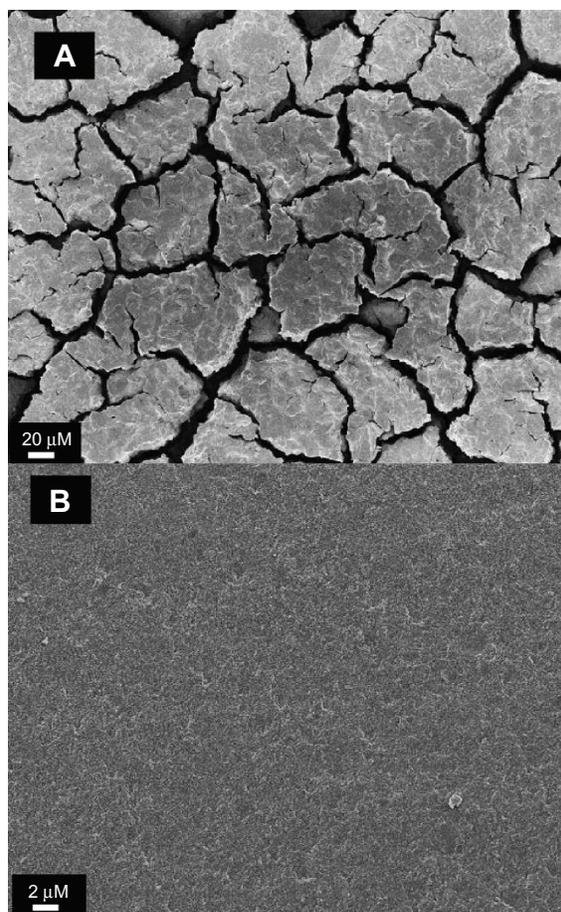


Figure 10 SEM micrographs of the GoHAP disk (**A**) and the commercial NanoXIM (Fluidinova, Moreira da Maia, Portugal) disk (**B**).

Abbreviation: SEM, scanning electron microscopy.

A biodegradation test clearly showed that GoHAP before the thermal transformation had a rate of degradation five times as large as NanoXIM. This test proved the unique properties and strong potential for future applications of this nanopowder. Additionally, the biodegradation study showed that during the first thermal transformation, GoHAP lost its unique properties and became similar in its rate of biodegradation to commercially available HAP nanopowder.

Biocompatibility study in vitro

Osteoblasts morphology

At 24 hours after seeding, the morphology of the cells on the culture plastic of the wells with both nanohydroxyapatite samples were comparable with the control, which was culture plastic without sample (TCP). Both correctly showed flattened cells with a phenotype typical for osteoblasts (data not shown). After confirming the adherence of the cells to the plastic, the medium was changed.

At 24 hours after changing to an osteoblastic medium, there was no difference between the morphology of the cells on the surface of the culture plastic next to the tested samples and the morphology of the control cells on the culture plastic without tested samples (Figure 11A and B). It was possible to observe a difference after fluorescent staining of the cells on the samples, with FDA staining the living cells (green color) and PI staining the dead cells (red color) (Figure 11C and D). No vacuolization or cell membrane lysis was found on the surface of any of the tested materials and only single dead cells were detected. Live cells were observed on both samples but they differed in their level of adhesion to the surface of the sample. On NanoXIM (Figure 11C) most of the cells were oval and poorly flattened, while on GoHAP (Figure 11D) the cells were similar in their morphology to those on culture plastic; they were flattened and seemed to adhere firmly to the surface of the sample.

On day 7, a significantly higher cell concentration was visible on the surface of the culture plastic around GoHAP compared with the wells with NanoXIM, although the cell morphology was similar to the control in both cases (Figure 11E and F). However, the cells visualized on the surface of the biomaterials at the same point in time had a substantially different morphology. GoHAP was covered by cells over its entire surface that were evenly adhered to the material and built a confluent layer (Figure 11H). The cells had a spindle lengthened shape, and only a few dead cells were found. In contrast, not all of the cells on the surface of NanoXIM showed the correct morphology (Figure 11G).

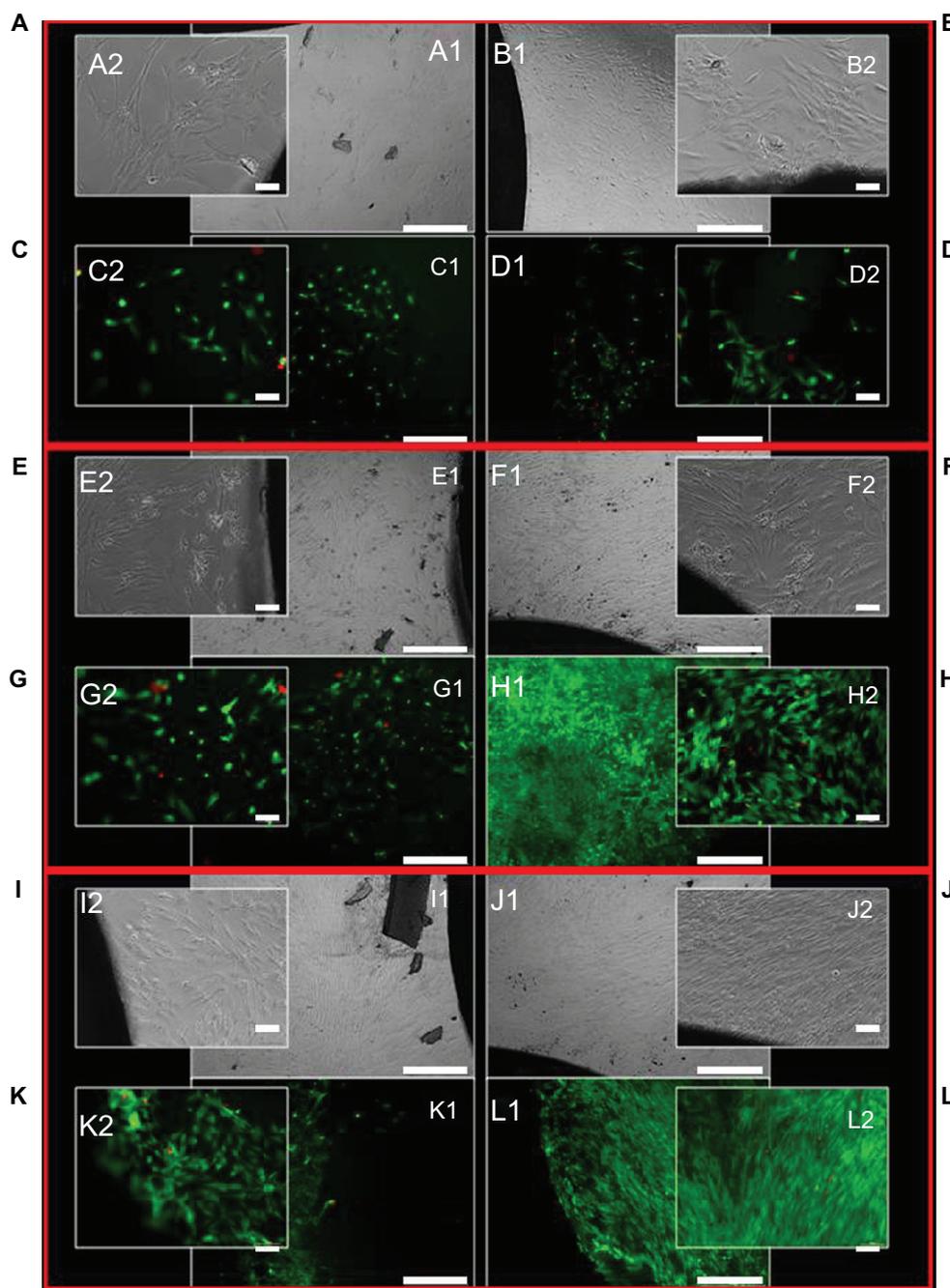


Figure 11 Microscopic images of in vitro cell culture on the nanohydroxyapatites surfaces. Images at 1 (A–D), 7 (E–H), and 14 (I–L) days after changing to an osteoblastic medium; 1 = 4× magnification (white scale bars 100 μm); 2 = 10× magnification (white scale bars 100 μm). (A, B, E, F, I, and J) Microscopic images from optical microscopy, showing morphology of cells on the surface of the culture plastic next to NanoXIM (Fluidinova, Moreira da Maia, Portugal). (A, E, and I) and GoHAP (B, F, and J) samples. (C, D, G, H, K, and L) Microscopic images from fluorescent microscopy showing morphology of cells on the surface of NanoXIM (C, G, and K) and GoHAP (D, H, and L) after FDA/PI staining.

Abbreviations: FDA, fluorescein diacetate; PI, propidium iodide.

The majority of the cells, although alive (FDA positive), were not flattened on the surface of the sample, and had a low confluence. The control cells (cultured on TCP) were correctly flattened, similar in shape to those visualized on GoHAP (data not shown).

On day 14, a multilayered culture was found on TCP near the GoHAP samples (Figure 11J), while single cells could still be easily identified near the NanoXIM samples. These latter samples showed a low stage of proliferation, but with the correct morphology (Figure 11I). The cells cultured on the surface

of GoHAP (Figure 11L), similar to those on culture plastic, displayed multilayer growth and were uniformly distributed on the surface of the samples. This observation was confirmed by scanning microcopy (Figure 12A and B). In contrast, on NanoXIM, the cells only partly covered the surface, although only the single cells were dead (PI positive) (Figure 11K). Even though there were significantly fewer cells compared with the GoHAP or TCP control, they were flattened and seemed to adhere to the surface of the sample (Figure 11K).

Quantitative assays in a culture on the biomaterials' surface

Figure 13 shows the results of the XTT assay and DNA measurements in the population of cells cultured in 96-well culture plates on the surface of the GoHAP and reference NanoXIM samples. The viability assay revealed a population of living cells growing in time on GoHAP (Figure 13A), which correlated well with the increasing amount of DNA (Figure 13B). The absolute values were lower compared to the TCP control at all time points: 21% on day 2, 34% on day 7, and 61% on day 14 in the XTT assay; and 113% on day 2, 56% on day 7, and 60% on day 14 in the PicoGreen test. At the same time, all of the results obtained for GoHAP were substantially higher compared to the results for NanoXIM, which served as a reference nanohydroxyapatite. In this case, the highest values in both the XTT and PicoGreen assays were obtained on Day 14 and did not exceed 5% and 4% of the TCP control, respectively.

Biocompatibility examination in an extract from the tested biomaterials

Figure 14 presents the results of the standard cytocompatibility study based on the cells' response toward the materials' extracts.

It was found that the behavior of the two tested powders (ie, GoHAP and NanoXIM) in contact with the culture medium was different (Figure 14H). NanoXIM immediately formed a homogeneous, nontransparent solution. In contrast, GoHAP instantly formed two layers: a constant nontransparent layer at the bottom of the extraction vessel and an almost clear supernatant above (Figure 14H). Following the preparation procedure as described in the Materials and methods section, the transparent extracts from both materials were obtained (Figure 14H) and used in the experiments.

The results of both the XTT viability assay and DNA measurements confirmed the expected cell response toward positive (triton X) and negative (corundum) controls (Figure 14F and G). HBDC viability as well as the DNA concentration in the extracts from both the investigated materials (GoHAP and the reference NanoXIM) were about 80% of the control's values in all the applied concentrations but one (ie, 100%). In that case both values were diminished. However, the activity and DNA concentrations were significantly higher for GoHAP compared to NanoXIM. For GoHAP the XTT assay was equal to 46% and the DNA concentration was 56%, compared to 27% and 46% for NanoXIM, respectively (Figure 14A and B).

The cell morphology was comparable to the control (cells in standard culture medium) in all investigated cases with the exception of two cultures: the culture exposed to triton X, independently of the concentration, and the culture exposed to 100% NanoXIM extract. In cultures exposed to triton X almost all of the cells looked dead. In the 100% NanoXIM extract case, about 30% of the cells were found to be dead (Figure 14C). The cell morphologies in 100% extracts of all types are shown in Figure 14C–G.

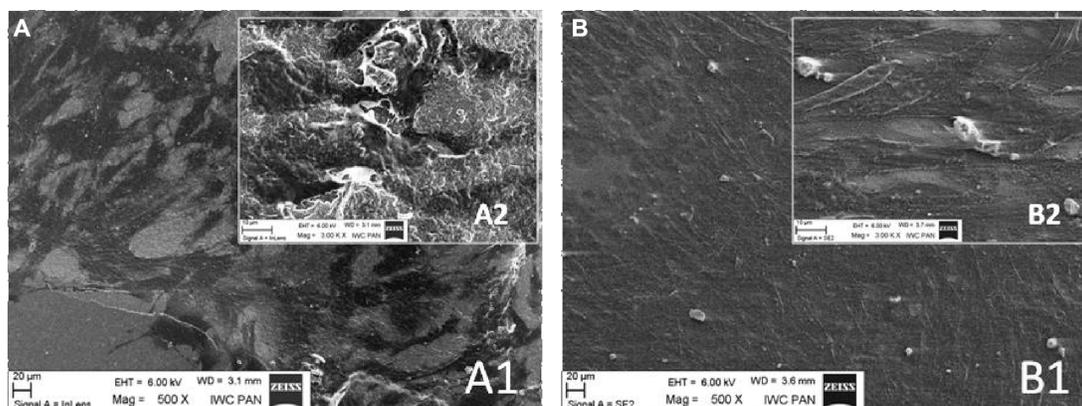


Figure 12 SEM micrographs of in vitro cell culture on the nanohydroxyapatites surfaces 14 days after changing to osteoblastic medium. Morphology of cells on the surface of (A) NanoXIM (Fluidinova, Moreira da Maia, Portugal) and (B) GoHAP. 1 = 500× magnification (black scale bars 20 µm); 2 = 3000× (black scale bars 10 µm). **Abbreviation:** SEM, scanning electron microscopy.

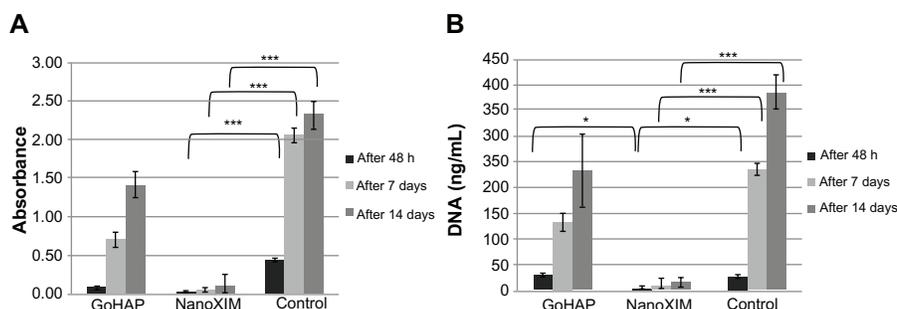


Figure 13 Representative results of (A) XTT assay and (B) DNA assay for GoHAP, NanoXIM (Fluidinova, Moreira da Maia, Portugal), and TCP (control, plastic surface) at three different time points (1, 7, and 14 days after changing to osteoblastic medium).

Note: Data presented as mean \pm SD.

Abbreviations: TCP, tissue-culture polystyrene; SD, standard deviation; XTT, 2,3-bis-(2-methoxy-4-nitro-5-sulfophenyl)-2H-tetrazolium-5-carboxanilide.

Discussion

By implementing a microwave heating technology with very high energy density, a method of HAp nanopowder synthesis was developed, which resulted in a material that mimics natural bone apatite. It was a pure, hexagonal non-stoichiometric HAp with a plate-like morphology, an average grain size of 6 nm, and SSA over 200 m²/g. As this powder had both a small particle

size much lower than 10 nm and a nonstoichiometric Ca/P molar ratio equal to 1.57, it achieved a high degradation rate and high biocompatibility. It was visible during the degradation test conducted for GoHAP and NanoXIM. GoHAP with a grain size of 6 nm and a Ca/P molar ratio of 1.57 had almost 5 times larger solubility than commercial nanopowder with an average grain size of 17 nm and a Ca/P molar ratio of 1.67. Furthermore, the

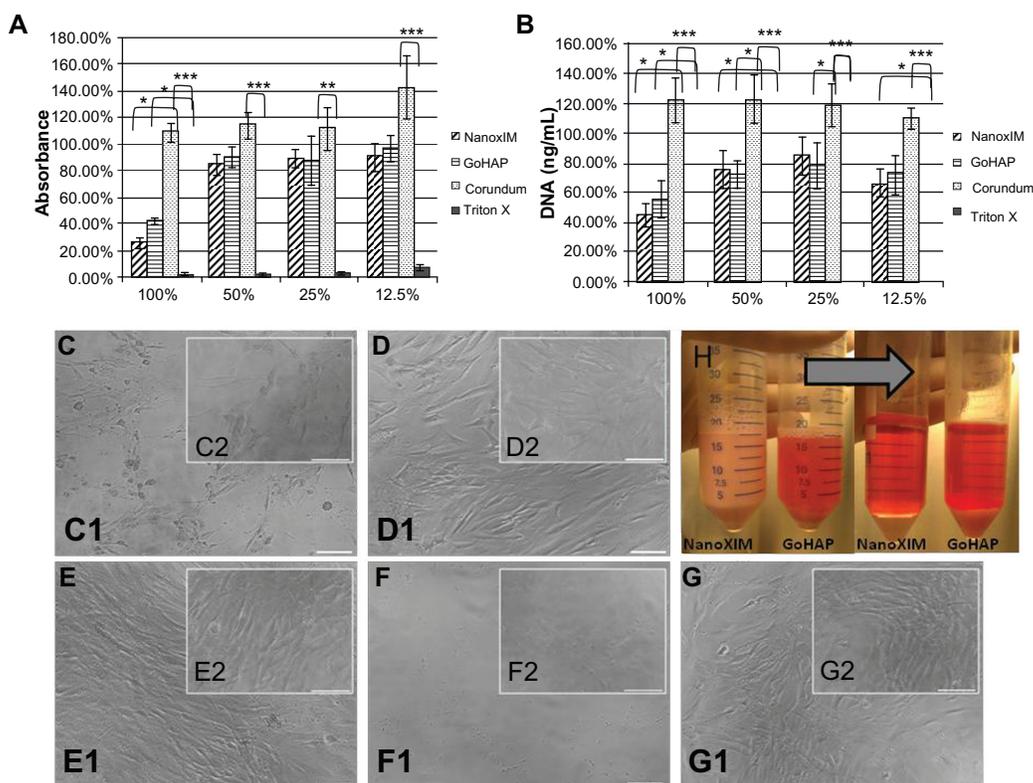


Figure 14 Biocompatibility examination in an extract of tested biomaterials. Representative results of (A) XTT assay and (B) DNA assay in four different dilutions of extracts (100%, 50%, 25%, 12.5%, and 6.25%) for both nanohydroxyapatites and positive (triton X) and negative (Al_2O_3 , corundum) control. Results compared with the viability of cells cultured in medium, which was defined as 100%, and presented as mean \pm SD. (C–G) Morphology evaluation of cells cultured in the presence of extracts from nanopowders at a concentration of 50 mg/mL of suspension 100% of extract. Optical microscopic images representative photos: 1 = 4 \times (white scale bars 100 μm); 2 = 10 \times (white scale bars 100 μm). (C) NanoXIM (Fluidinova, Moreira da Maia, Portugal); (D) GoHAP; (E) corundum; (F) triton X; (G) TCP (control, plastic surface); (H) extracts from nanopowders (GoHAP, NanoXIM) before and after centrifugation.

Abbreviations: SD, standard deviation; TCP, tissue-culture polystyrene; XTT, 2,3-bis-(2-methoxy-4-nitro-5-sulfophenyl)-2H-tetrazolium-5-carboxanilide.

literature shows that osteoblast adhesion is strongly driven by the material surface size: a decrease in particle size increases the osteoblast adhesion which is directly related to the larger surface area, seen in this paper by comparing NanoXIM (with a surface close to 120 m²/g) to GoHAP (with an SSA almost twice as large, equal to 236 m²/g in in vitro studies).¹² At all time points the cell culture on GoHAP had superior results, showing a multilayer cell culture with correct morphology over the entire material surface. Annealing of GoHAP powder at a temperature of 375°C decreased significantly the degradation rate. It seems that the advantages for the possible application of GoHAP nanoparticles for the development of resorbable bone scaffolds are lost when the material undergoes the first thermal transformation at 375°C.

The application of microwaves enabled a precise control of the crucial process parameters: time and pressure, and ensured high purity of the reaction conditions. Compared to conventional heating, microwaves transmit energy directly to the entire volume of reaction media. The rapid heating process leads to overheating of the solution and precipitation of nanocrystals with a nonequilibrium structure, similar to those produced in the human body, and reduces temperature gradients, which allows us to obtain a uniform nano-grain size distribution. The process is as short as 90 seconds and subsequently the solution is rapidly cooled, so that no secondary grain growth takes place.^{18,39}

Conclusion

A highly biocompatible HAP nanopowder, which we called GoHAP, was successfully synthesized using a novel method in which heating was accomplished with high density microwave radiation in a closed vessel. This allowed for a unique material to be produced inexpensively with an easily scalable method. The obtained nanopowder had unique properties; it was a pure, fully crystalline HAP nanopowder with a hexagonal structure. Its grains had needle-like morphology, an SSA close to 240 m²/g, a Ca/P molar ratio of 1.57, and an average particle size (estimated using TEM) under 6 nm. As the degradation study demonstrated in conditions simulating the human body, its solubility was 5 times larger than that of commercial nanopowder (NanoXIM). These properties were lost when the powder was heated to the thermal transformation temperature of 375°C, at which point its structure became dehydrated and the material's solubility matched that of the commercial nanopowder. GoHAP showed excellent biocompatibility properties, as confirmed by in vitro tests. Its properties were good not only as a surface for cell growth (which was demonstrated by FDA/PI staining and the XTT results) as no vacuolization or cell

membrane lysis was found on the surface, but it also promoted cell proliferation as confirmed by the results of a DNA assay. The cells were equally distributed along the available ground surface, building a single confluent layer. Furthermore, the cells presented a correctly flattened phenotype, maintaining a morphology typical for bone cells.

The presented results were achieved due to an extremely short synthesis time with high energy delivered to the reactants using a unique microwave reactor. With its unique features, the obtained GoHAP nanopowder is a promising material for resorbable bone implant fabrication.

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Disclosure

The authors report no conflicts of interest in this work.

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